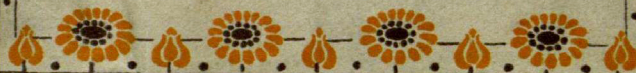
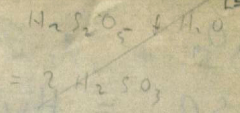


THEYER & HARDTMUTH
WIEN · I · KÄRNTNERSTRASSE · 9

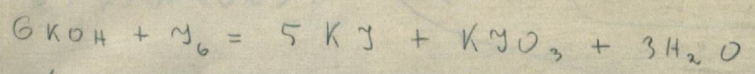
Hydrogene und
anorganische Chemie
Kraus. V. 1. Teil



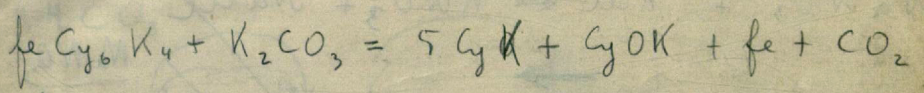
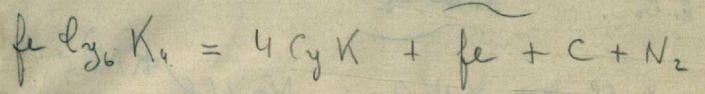
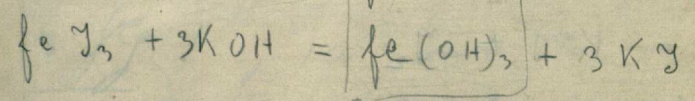
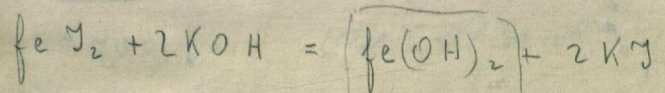
W.



7/3 ~ 2/2 = 2/2

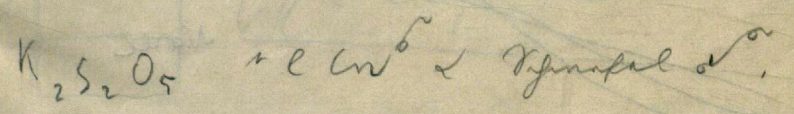


1/2 ~ 2/2, e 2/2/2

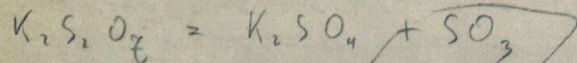
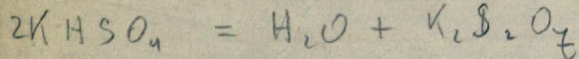


1/2 Ba 2/2 R

K all 2/2, 3/2 ~ 2/2
 KOH



K_2SO_4 is soft mass, very soft.



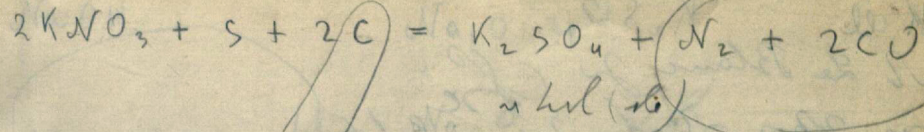
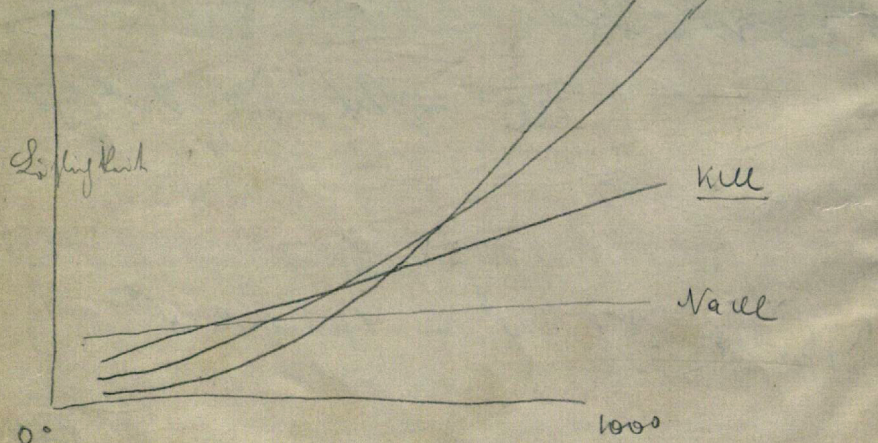
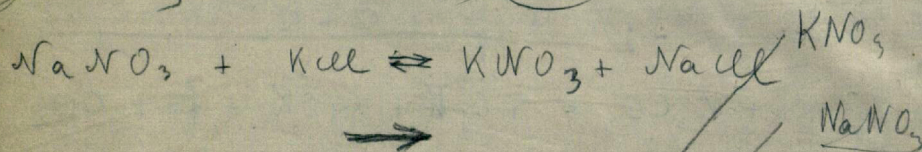
CuO

he is 9 = key of books 2.

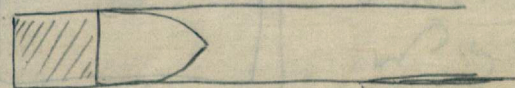
$KMnO_4$ in red



from red = blue.

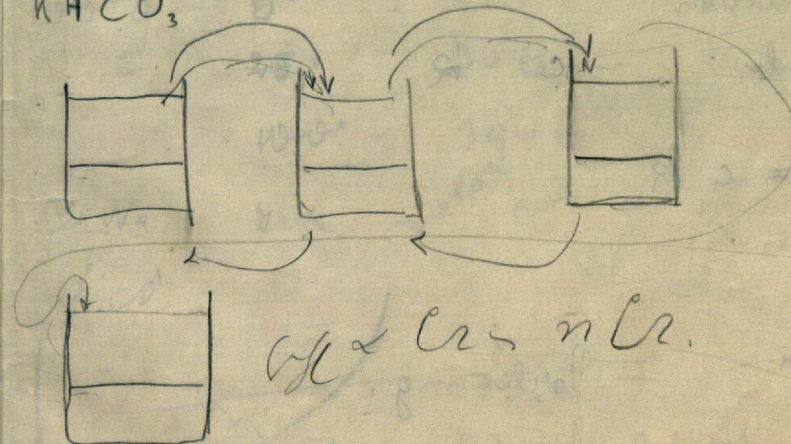
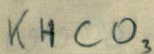
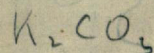


2 ~ P.H.S.O.N.S.



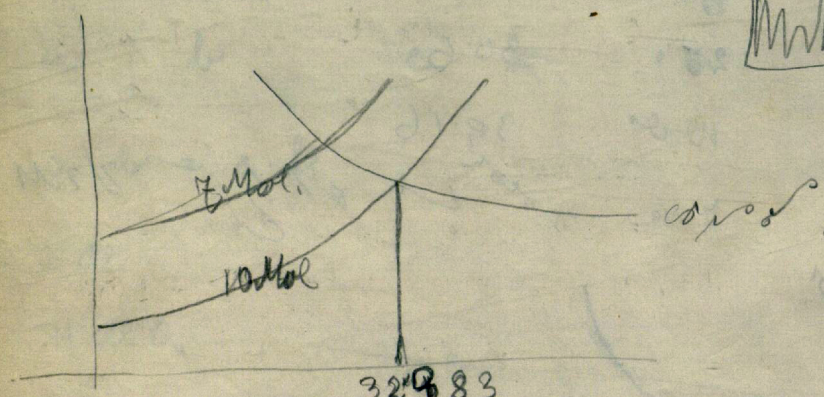
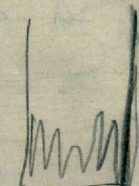
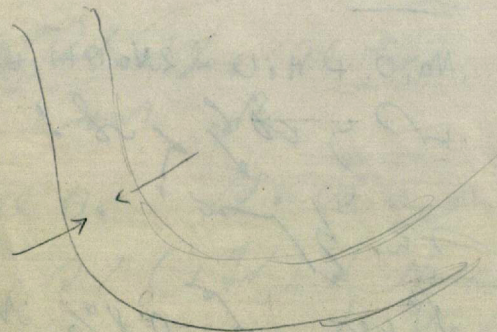
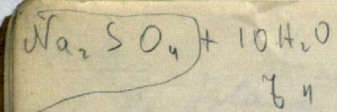
~ 700, p. 1022 (David)

2 sp! 100 ~ 100 / 2.

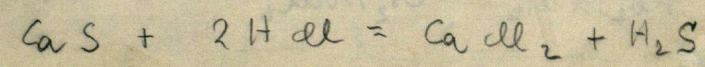
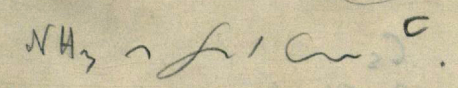
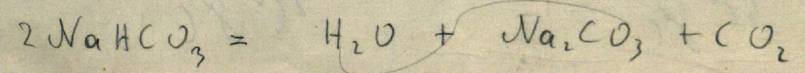
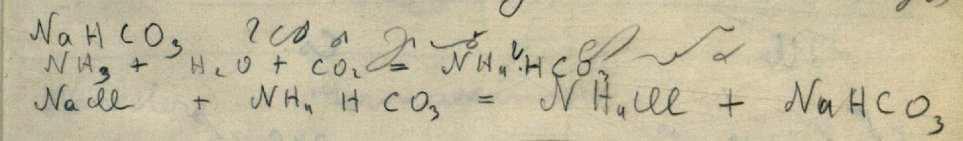
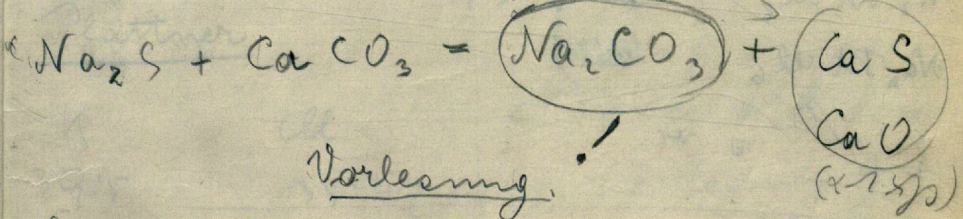
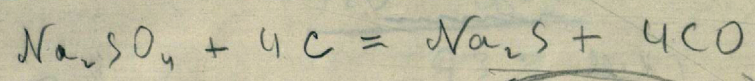
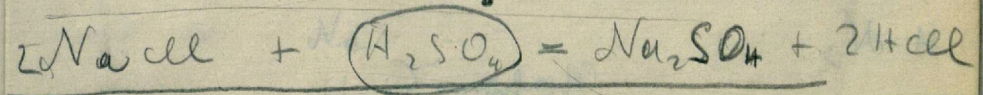
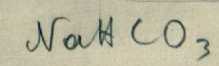
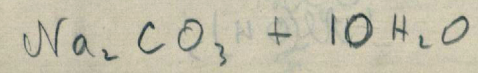
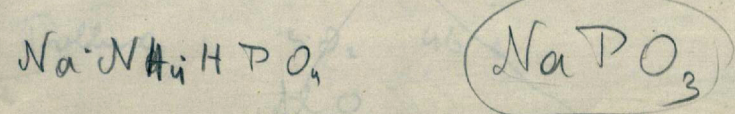
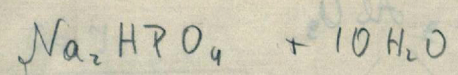
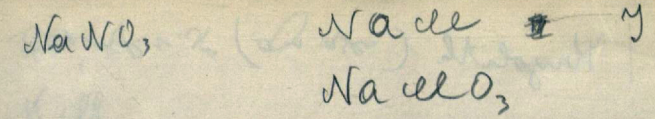


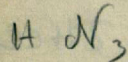
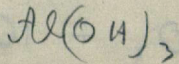
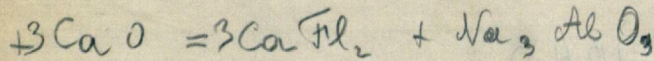
off 100 ~ 100.

Natur

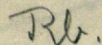
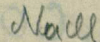
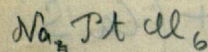


Handwritten notes in German, including "Vergleich" and "Zusammenhang".



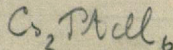
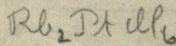
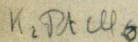


Wongjunt:

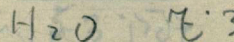
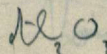
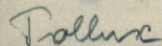
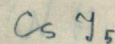
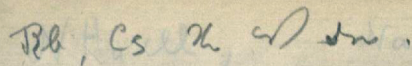


380

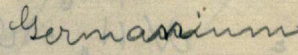
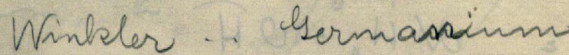
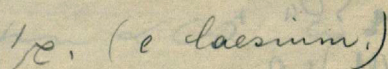
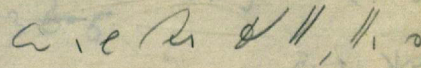
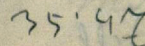
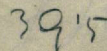
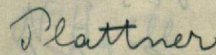
270 J.C.



Tatrate

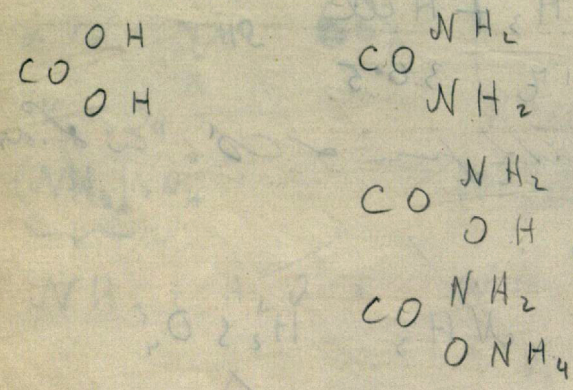


93 %



all NH_4 ✓ ✓ all Na ✓ all K ✓
 20 ✓ $BrNH_4$ YNH_4

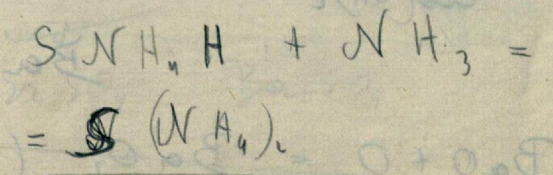
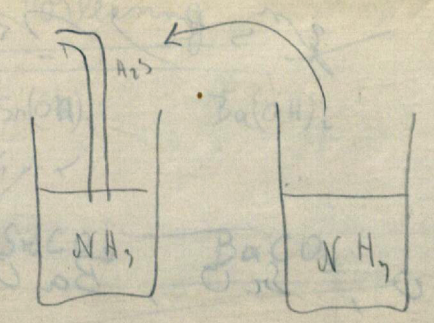
$BrNH_4$ 7. 20 1 20
 20 20, $(NH_4)_2CO_3$
 ✓ $CaCO_3$ & NH_4Cl
 20 20 ~ $(NH_4)_2CO_3$
 a NH_4HCO_3
 $CO_2(NH_3)_2$



all NH_4 ✓
 20 20

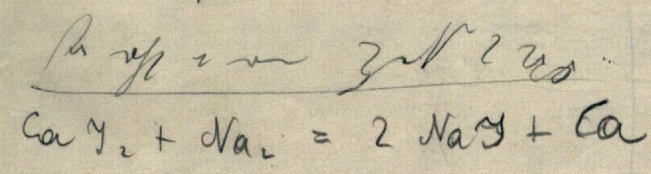
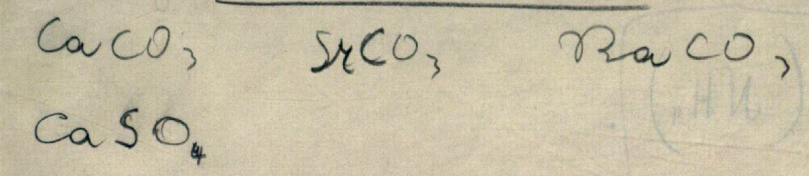
20 20
 $CO_2(NH_3)_2$

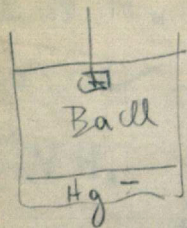
$3(NH_4)_2$
 $5NH_4H$



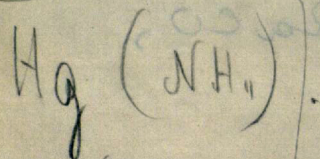
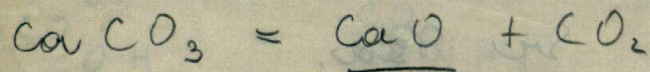
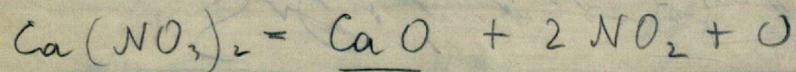
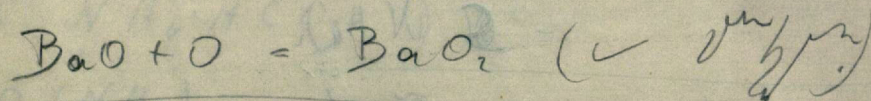
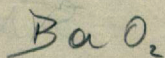
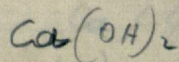
NH_3 2 20 20 20 20

Ca SO Ba.



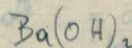
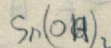
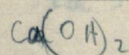


sl. CaO , SrO , BaO

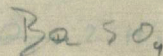
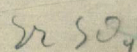
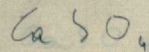
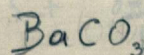
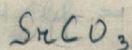
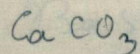


N/2.

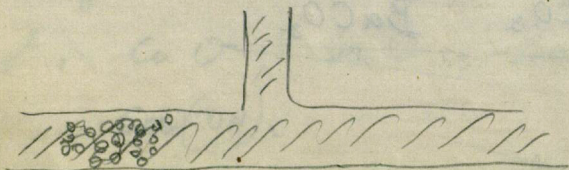
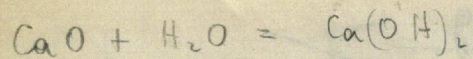
Vorlesung



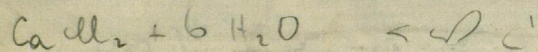
sl. ✓



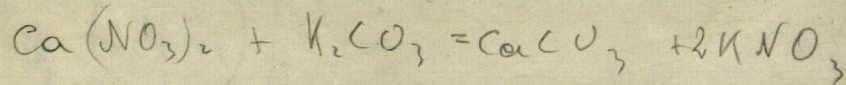
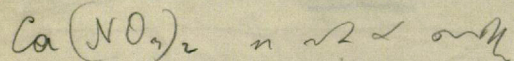
547°



Calc. n. 1. in 100 g. of 100 g. of 100 g.



100 g. of CaCl_2 (-400).

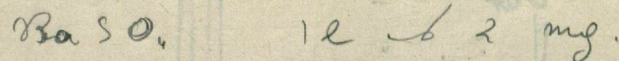


Calcite, n. 1. } 100 g.

Orthogonit n. 2. }

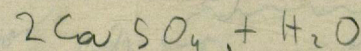
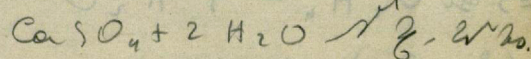
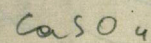
Orthogonit n. 2. }

4. 5. Calc. n. 1. in 100 g.

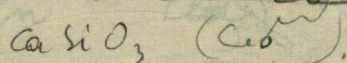


BaS sp. 1. in 100 g.

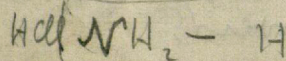
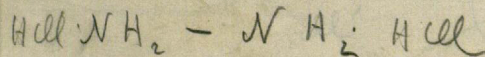
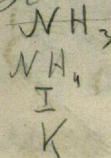
2. 2. in 100 g.



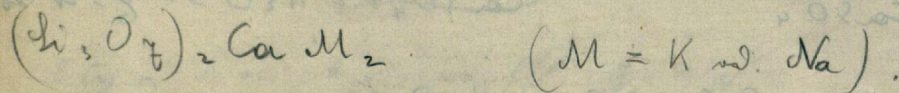
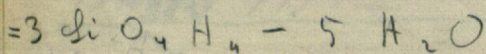
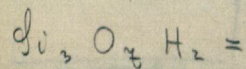
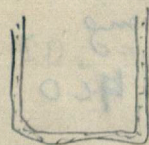
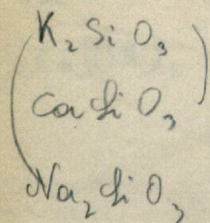
Worlesing



up 100 g.
Ba n.
up 100 g.

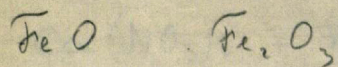
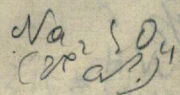
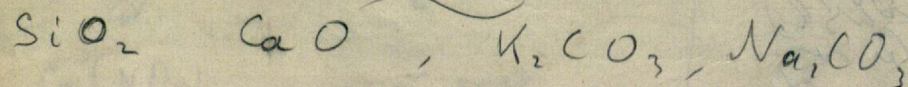


flüssig.



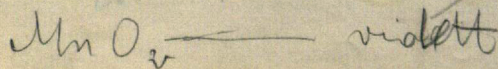
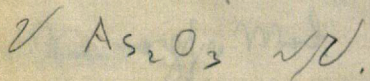
grünlich

weiß, d. H. gelb

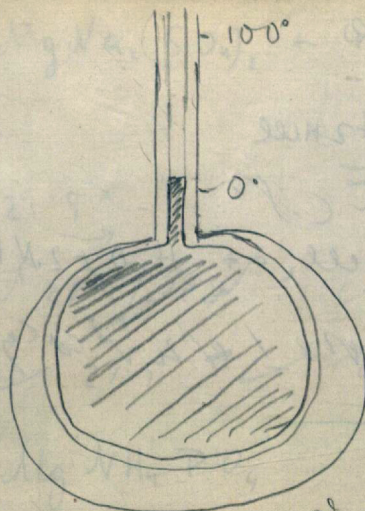


grün

gelb - rotbraun
weiß



weiß

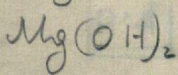
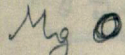
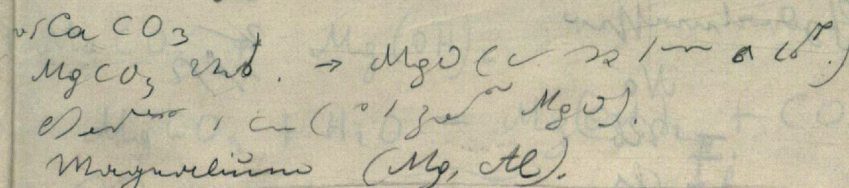


0°
250°

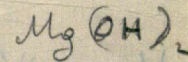
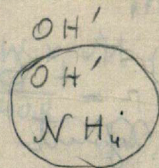
Vorlesung

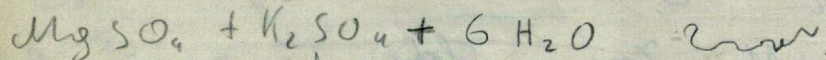
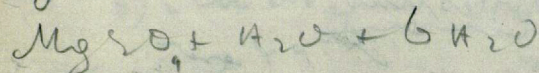
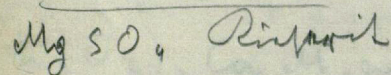
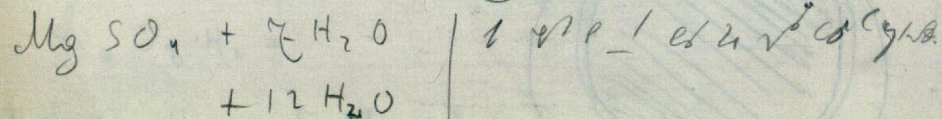
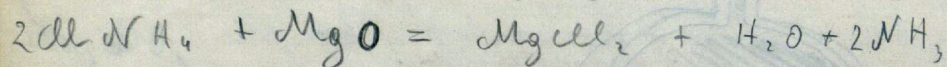
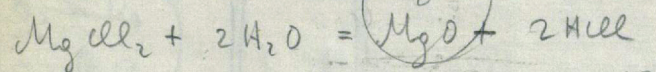
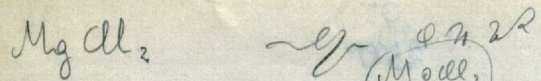
Be, Mg, Zn

old.



all





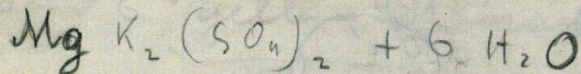
gulfondurururur.

Fe Na

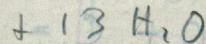
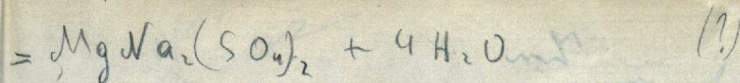
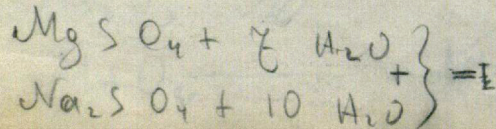
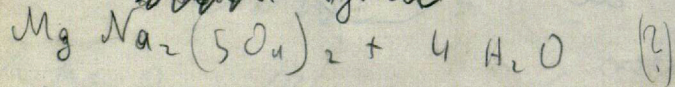
Zn Rb

Cu Cs

NH₄

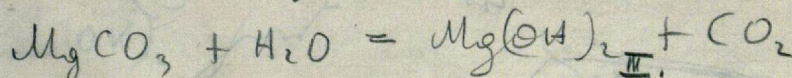
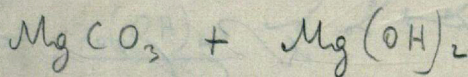
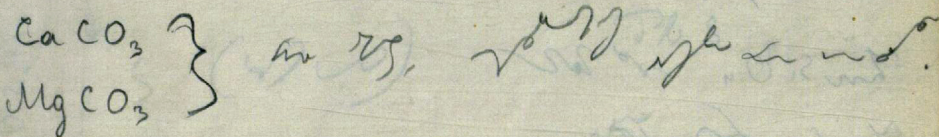
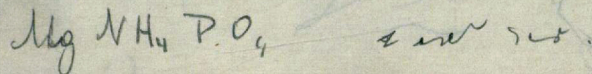


~~gulfondurururur.~~ Rinsed



21.5°

21° 1/2° 1/2° 21° 1/2° 21° 1/2° (23)



Be 13 Beryllium

9.4

II.

Mg

$$\text{S.W} \times \text{M} \div 6.4$$

$$\frac{6.4}{0.408} = 13$$

44 Rb

$$\text{M} = 29.9$$

$$35.5 \times 2$$

$$71.0$$

$$71.0$$

Zn

ZnO

ZnCO₃

ZnS



Vorlesung.

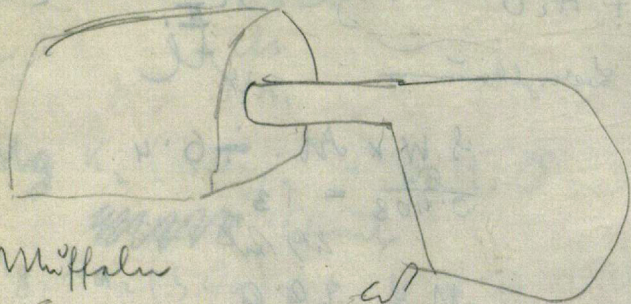
ZnO 1. f. g.

ZnCO₃ 2. f. g.

ZnSiO₄ 3. f. g.

ZnS 4. f. g.

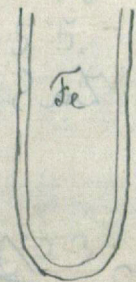
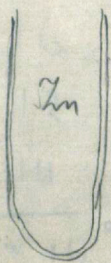
etc. etc. etc.



Muffel

f. g., Zn + ZnO.

12. 1. 12.



etc. etc. etc.

etc. etc. etc.

ZnO

↓

ZnS (etc.)

Zn(OH)₂ etc.

Zn(OK)₂ f. g.

ZnCl₂
ZnBr₂
ZnI₂

Zn OH

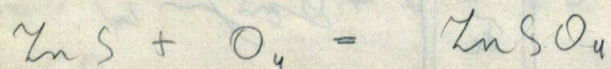
ZnSO₄ + 7H₂O

ZnSO₄ + H₂SO₄ + 6H₂O

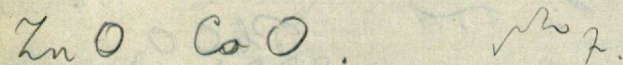
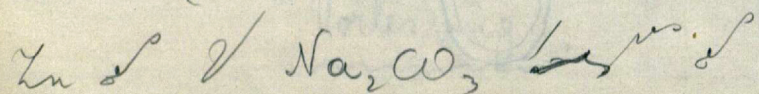
Vergleiche.

PbCO₃
↓
PbS

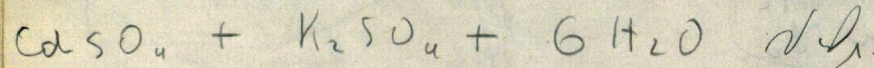
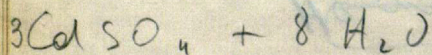
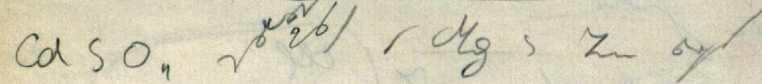
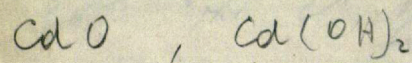
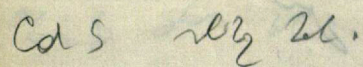
etc. etc.



cl 25 e -



Cd = 112



Cd relg 112

H 1 Na 23

O 16 Cl 35.5, 35.47

C 12

N 14

O = 16 Li = 7.03

Na = 23.05

S = 32.06

H = 1

O = 15.88

Diff.

O 16

- 16

S 32

- 47 16.3

Cl 35.5

- 48 16.3

Fe 127.7

Na 23

- 16

K 39

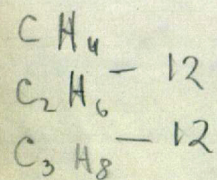
- 47 3.16

Rb 85

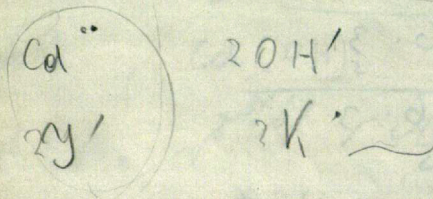
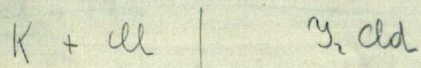
- 48 3.16

Cs 133

$$\begin{array}{r} \text{dl} \quad \text{Bn} \quad \text{y} \\ 35.45 \quad 79.96 \quad 126.85 \\ \underline{35.45} \\ 162.30 : 2 = 81.15 \end{array}$$



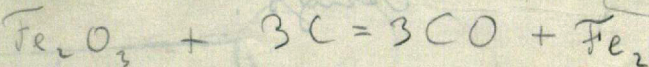
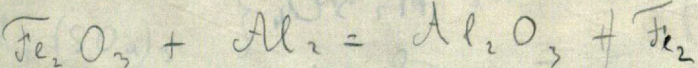
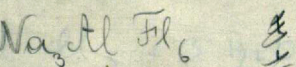
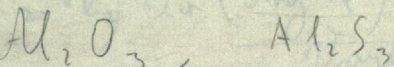
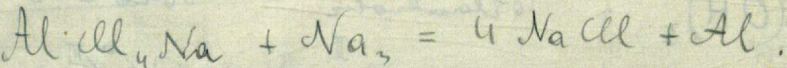
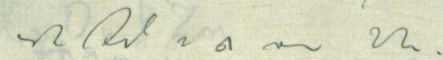
Vorl.



H=1	At. W.	Na = 23	23.7
Li = 7	5.9	Mg = 24	12.8
Be = 9	5.6	Al = 27	10.7
B = 11	4.1	Si = 28	11.2
C = 12	3.6	P = 31	15.5
N = 14		S = 32	15.7
O = 16		dl = 35	25.6
F = 19			

I	II	III	IV	V	VI	VII	VIII
Li	Be	B	C	N	O	F	
Na	Mg	Al	Si	P	S	dl	
K	Ca	Ti	V	Cr	Mn	Fe	Co Ni

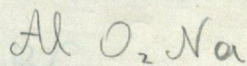
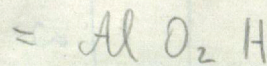
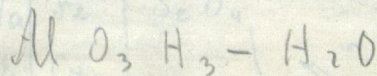
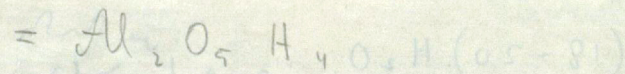
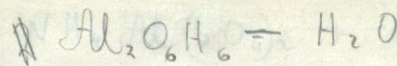
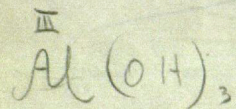
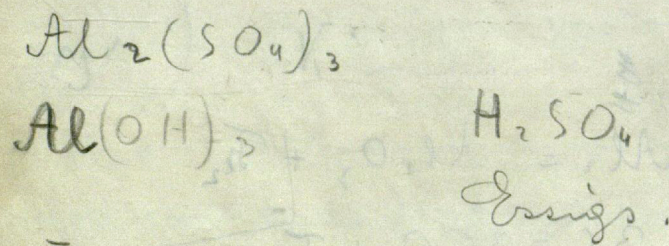
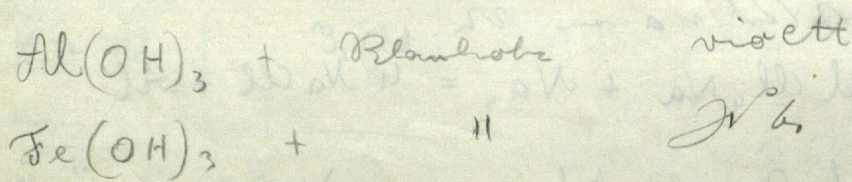
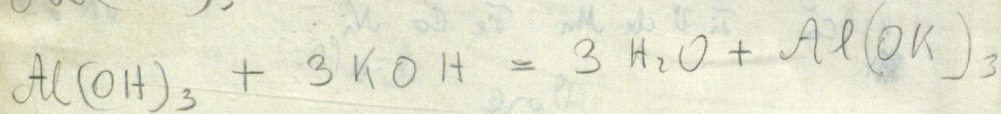
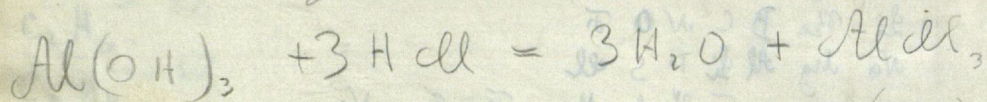
Vorl.



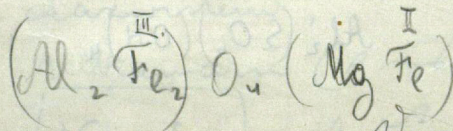
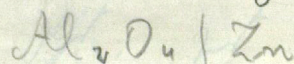
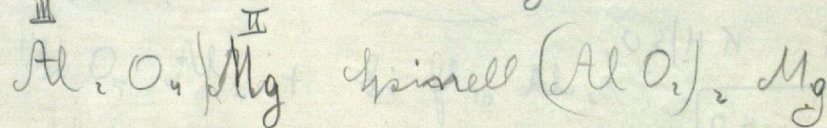
Germineur.

II
Al.
Al₂O₃ *Hydrat*
PbO *alkalisch* *von* *(OH)*.

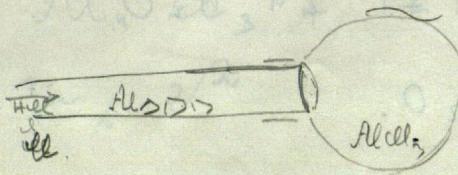
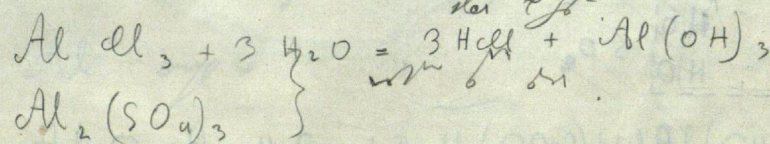
oRrR.
Al₂O₃ *aus* *PbO*.

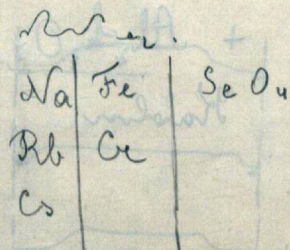
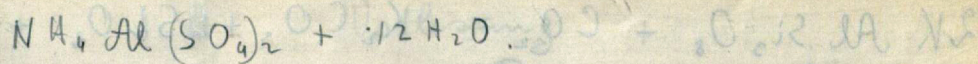
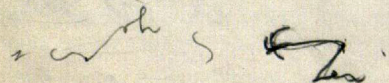
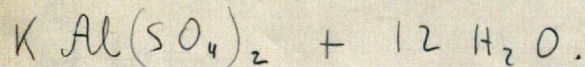
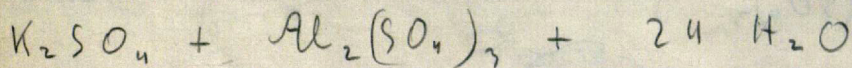
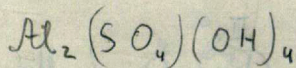
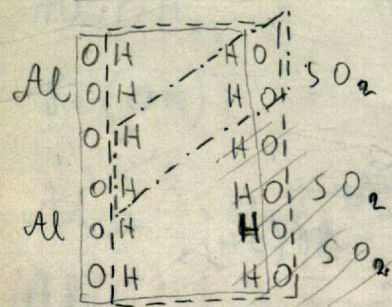
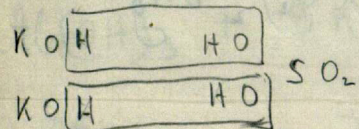
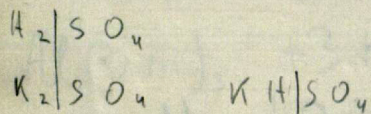
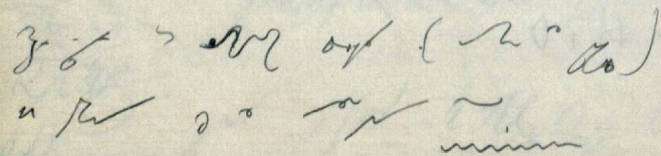
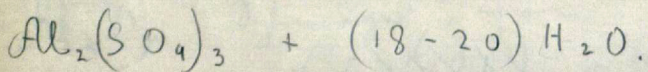


III

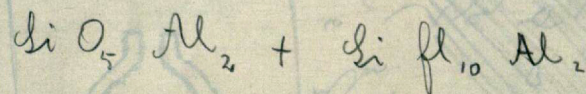


Vorlesung





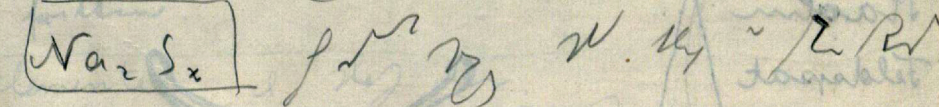
Alkalib.



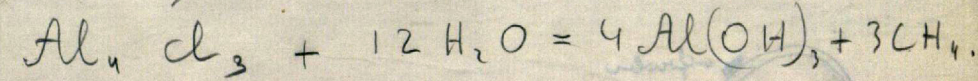
sept, 20.

Laserstein.

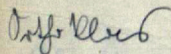
ultramarine (p. 10).



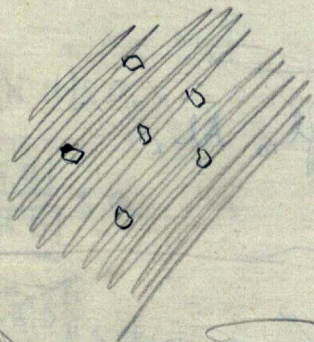
Al oxy.



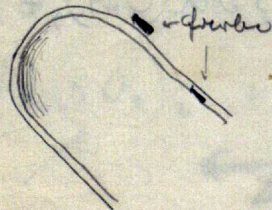
$\sqrt{2}$ $\sqrt{2}$.


$$+ \text{Al}_2\text{Si}_2\text{O}_7$$

137 (1) 1894



Maolin
Feldspat
Quarz



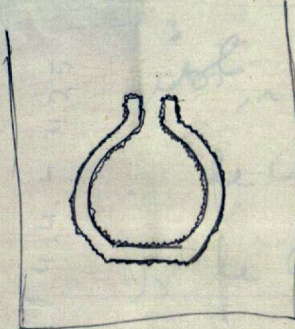
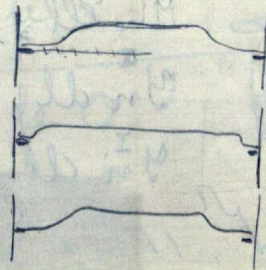
Vorlesung . muller

Porzellan

Reaktion in der gleichen

Steinzeng

2d & 2¹ ✓



20 Dec 19, 2000

Fayence

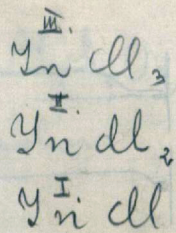
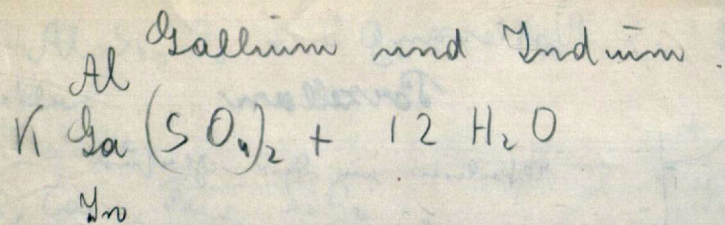
Bleinglaswe

Spillman.

clemente, a day or

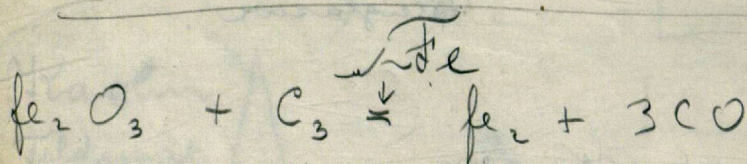
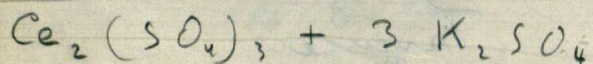
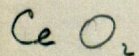


С. Р. Г. С. Р.

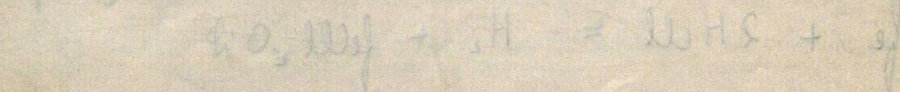
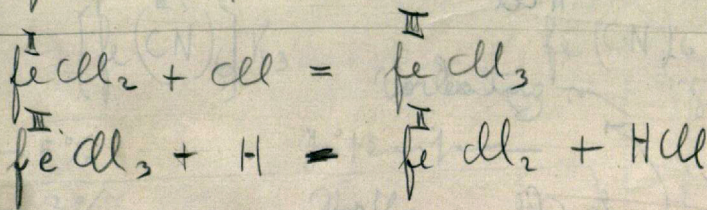
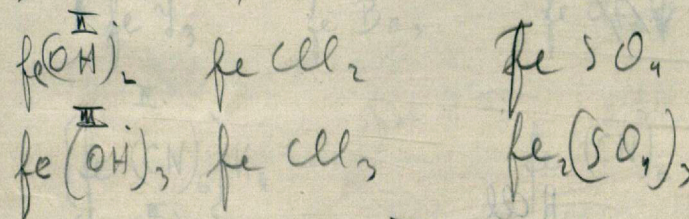
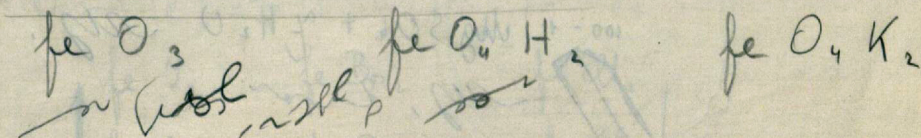
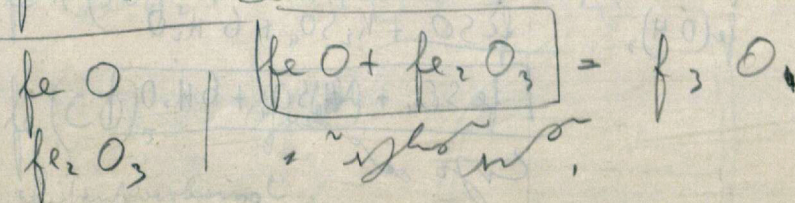
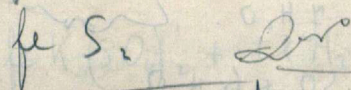
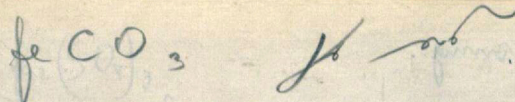
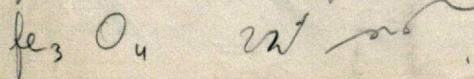
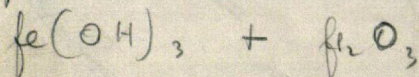
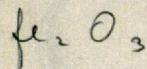


1. $100\% H_2O$ $100\% H_2O$

Gallium Indium.



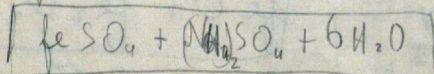
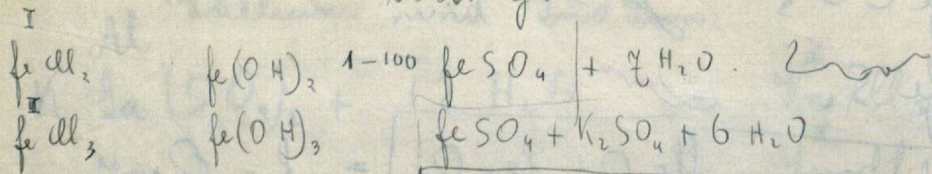
$1600^\circ F.T.$



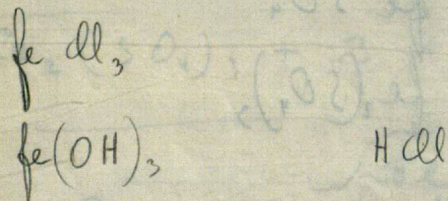
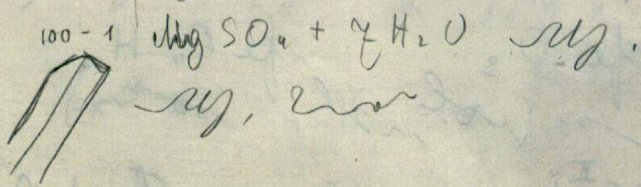
Gefüge

5.10

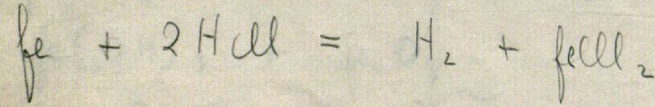
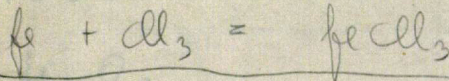
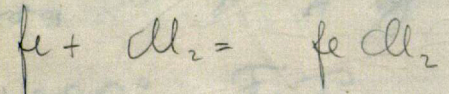
Vorlesung.



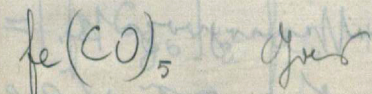
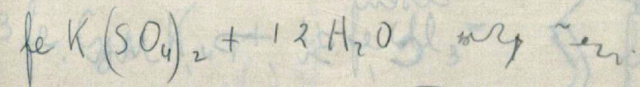
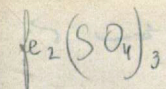
Ergo



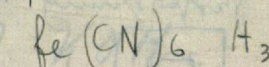
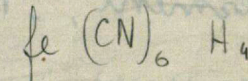
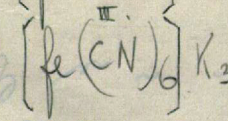
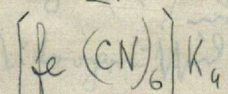
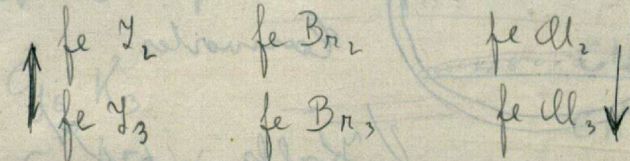
unlöslich in Wasser



x+2

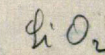
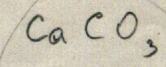
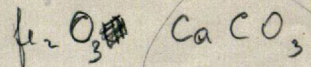


Feinprobe

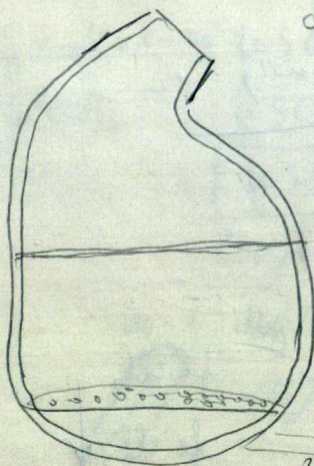


Vorlesung

- 5% ... Güte
- 3% ... Maß
- 1% ... Feinprobe



Gefüge



Eisen

C, Si, P, Mn.

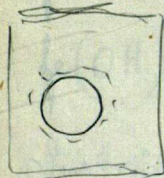
Schwefelwasserstoff.

200 g ...

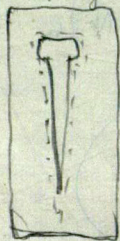
Carbonat.

Kalk

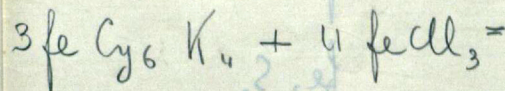
Thomasmehl, P, C, Si, Mn.



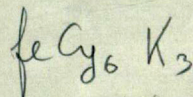
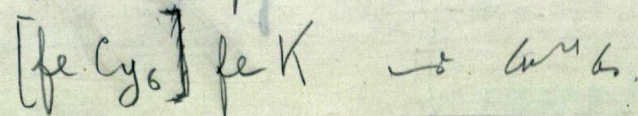
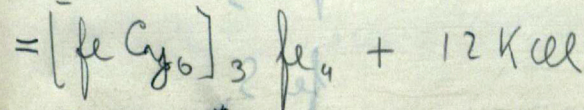
Verflüchtigung ...



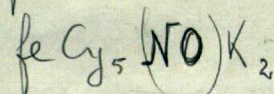
...



Fe^{III}



CNH

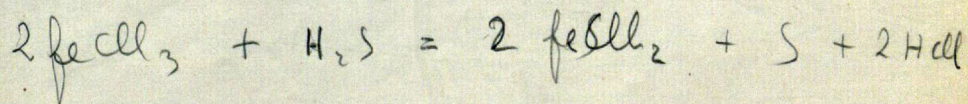
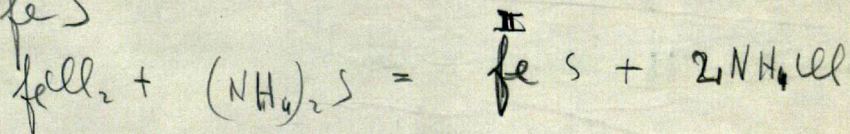


CNOH

CNSH

Nitrogenium und Kalium.

3 ... made 2 KOH

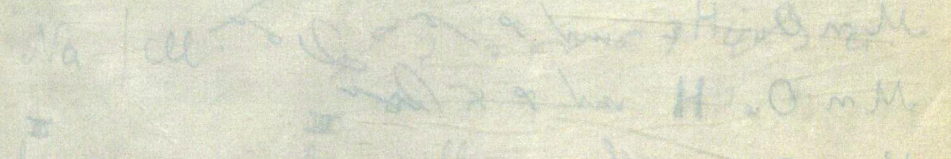
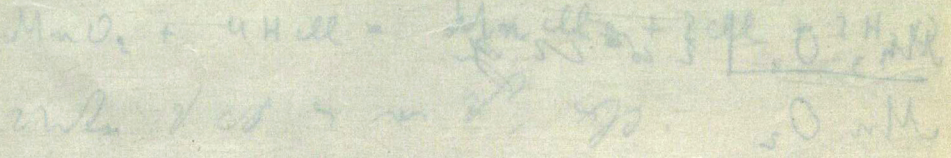
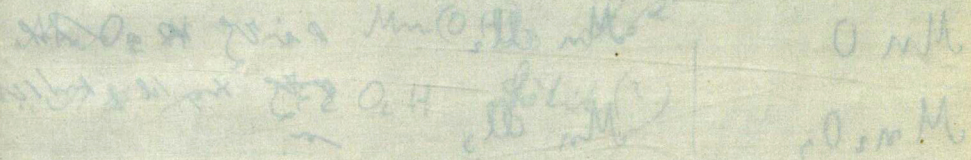
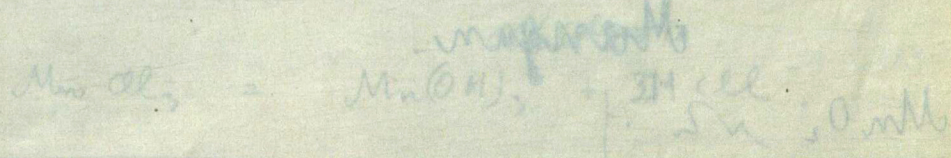
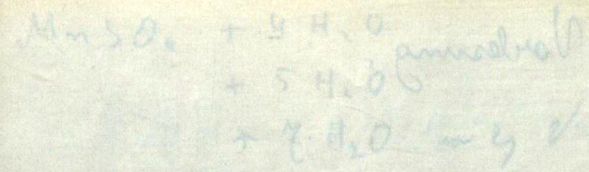
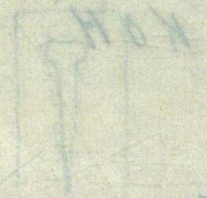
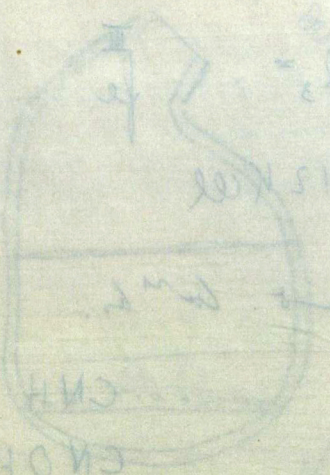


FeO
Fe₂O₃

FeS

Fe₂S₃

FeS₂



π
 $\text{Fe} \quad \text{K}_2(\text{C}_2\text{O}_4)_2$
 $\frac{1}{m}$
 $\text{Fe} \quad \text{K}_3(\text{C}_2\text{O}_4)_3$

MnO₂ ✓

Mn_2O_3	$MnCl_3$
-----------	----------

Mn₃O₄ | 2 6th 12 g.

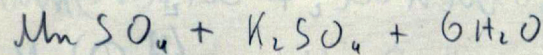
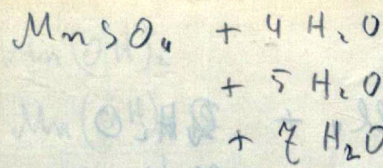
 MnO_2

$\text{MnO}_4 \text{ H}_2$ und H_2 $\rightarrow$ 2 or.

MnO₄ H and x x

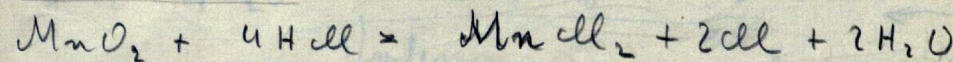
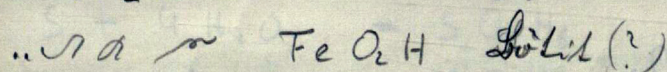
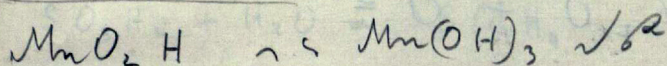
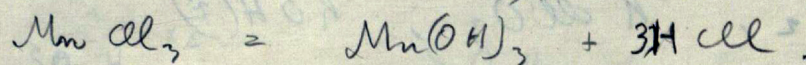
Mn_2O_7 solⁿ g.

Livingston
or x, he is in the

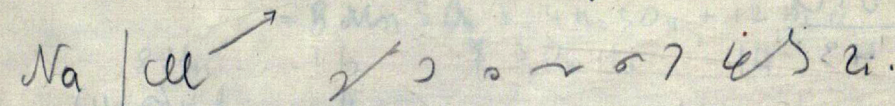


Aln_2O_3 p.p. o.l.v. Laminat.

class :



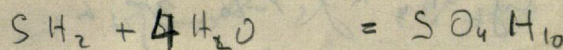
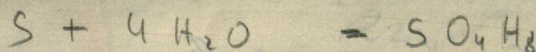
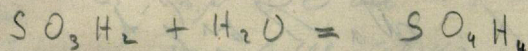
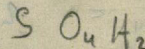
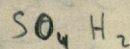
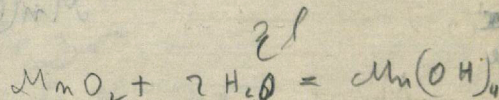
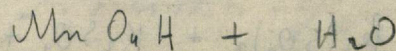
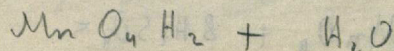
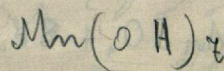
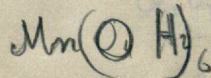
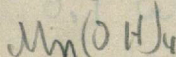
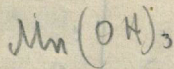
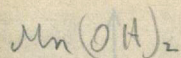
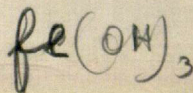
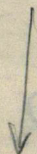
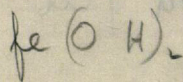
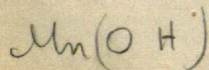
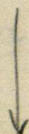
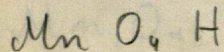
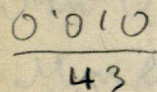
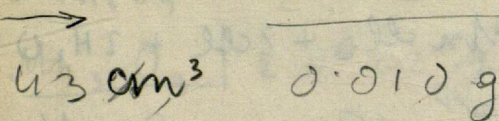
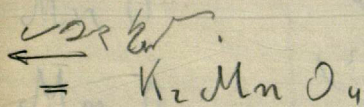
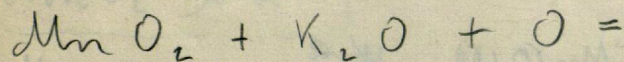
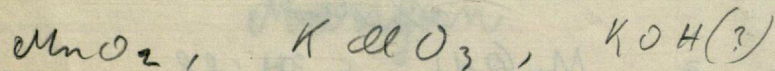
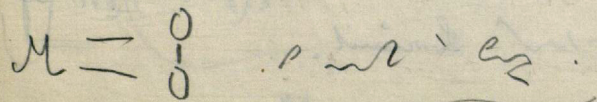
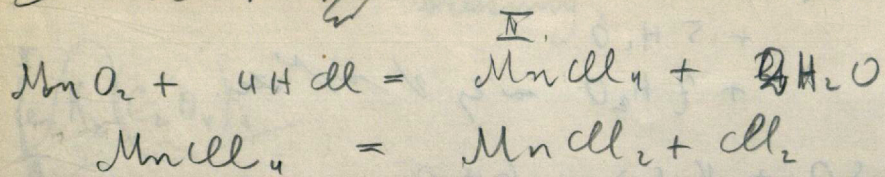
W. B. & S. W. B. G.



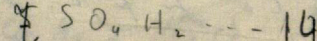
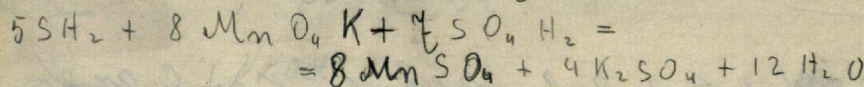
MnO_2 violett.

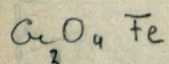
2.

✓ 1 re 21 6



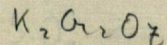
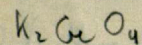
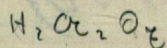
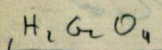
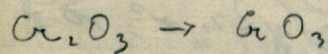
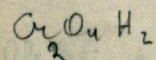
Verlesung





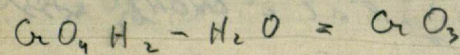
Ref

Ferrochromit



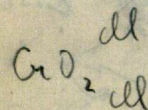
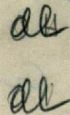
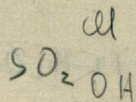
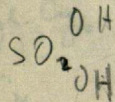
gelb

not

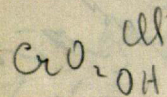


2. 4. 2011 5. 2011

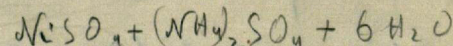
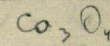
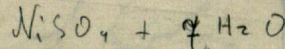
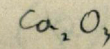
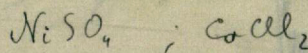
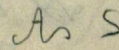
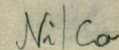
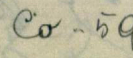
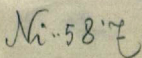
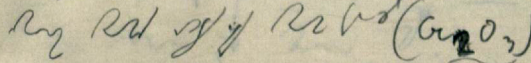
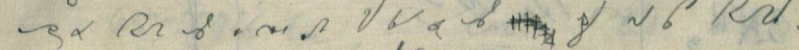
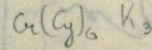
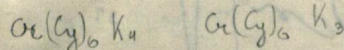
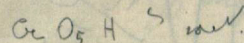
Vorlesung.



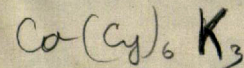
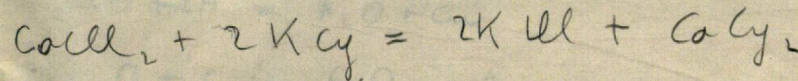
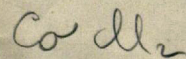
chromylchlorid

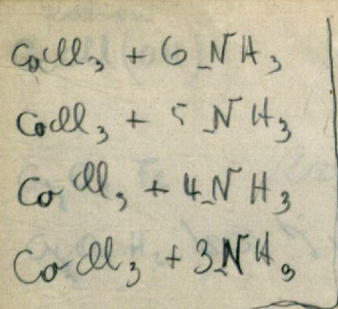


Chlorochromsäure

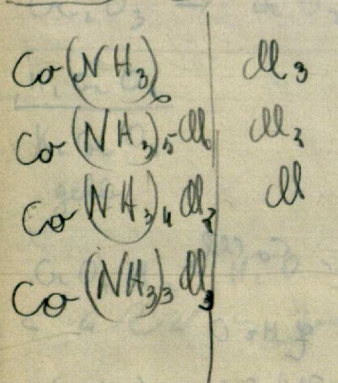


Die Hesperiden

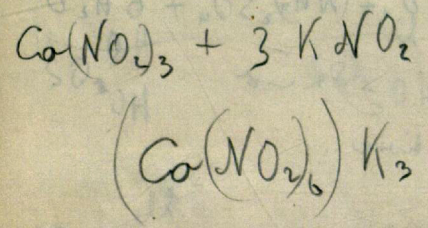




as in 1st Dec
2 in 2nd Dec
1 in 3rd Dec
- in 4th Dec

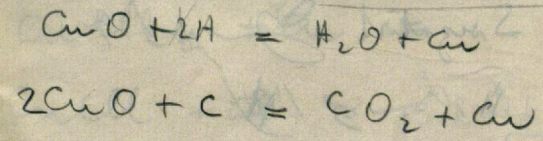
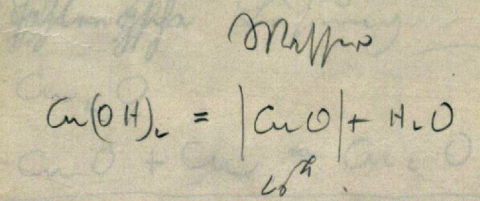


1st 25% to be
others 1st Dec

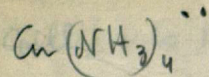


Vorlesung

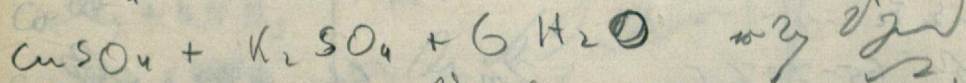
an ...
an ...
65-70% Cu, 2 Zn. an ...
V ...
V Mn ...
V Al ...
CuO, Cu₂O, Cu₄O ...
Cu₂Cl₂, CuCl₂
...



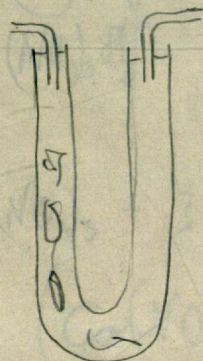
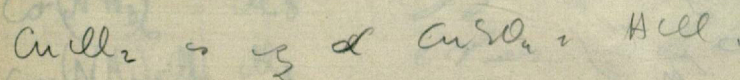
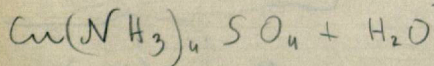
flamant ...



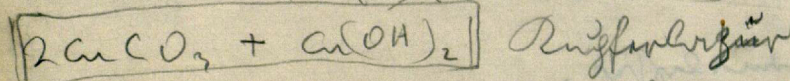
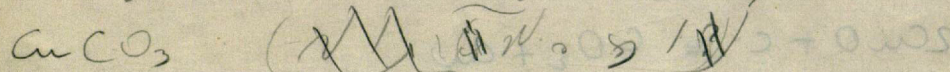
May 21



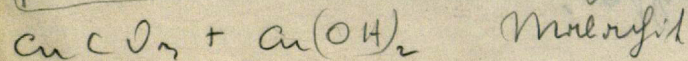
Vorlesung



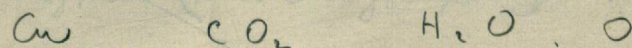
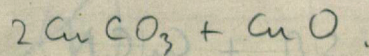
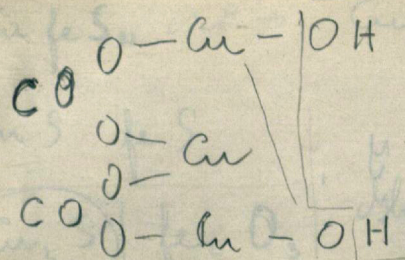
Ringarvorfund - - regelat 2 p 7



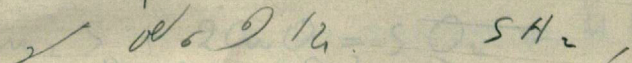
Ruffenberger



Melnyk

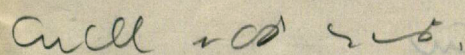
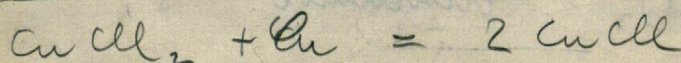
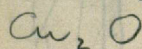


Wm. F. F. F.

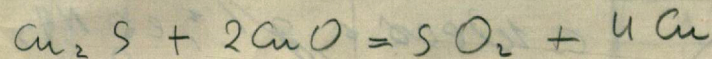
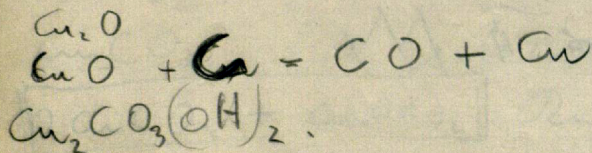


May 25, 1901.

Salmon for Living (Living) . . .



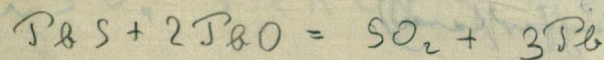
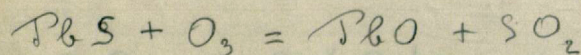
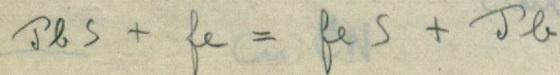
Lu.



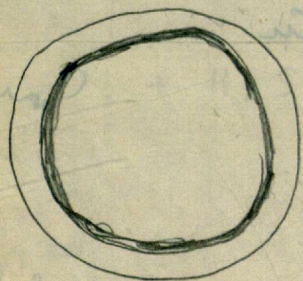
in Pb
♀ h

Vorlesung

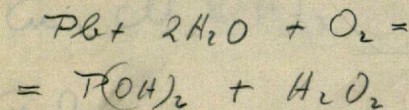
PbS a 2v.



← in, 1/2.



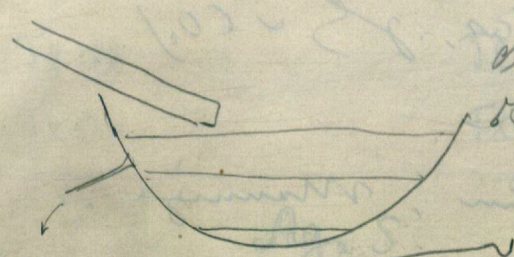
a Weisses Pulver, es NH_3
ist, es ist ein Pulver
E.



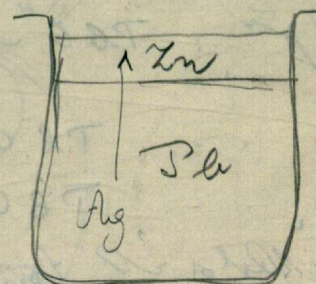
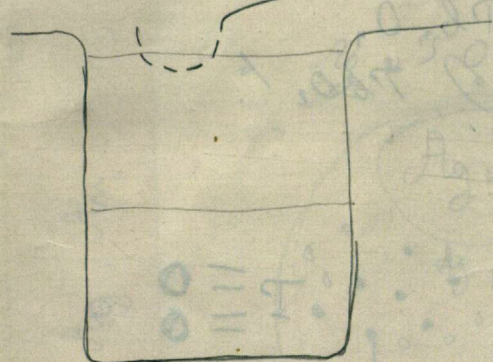
a 2v ~ 2v a 2v 2v

Pb₂O, PbO, PbO₂, Pb₃O₄, Pb₂O₃.

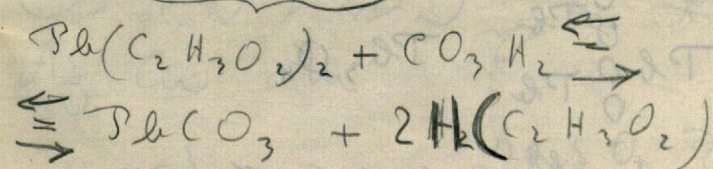
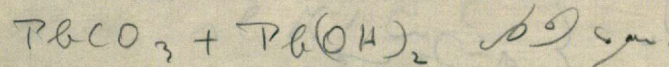
PbO d'ist a 2v d'ist a Pb ~ 1.
d'ist a 2v d'ist a 2v.

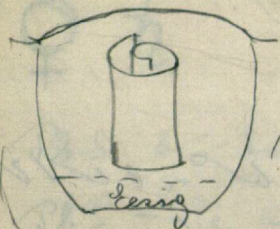


d'ist a 2v d'ist a 2v
d'ist a 2v d'ist a 2v



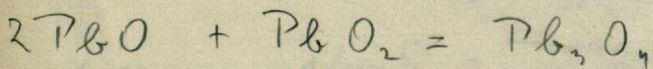
PbCO₃ d'ist a 2v





Whiff (ar. $2S \sim CO_2$)

Pb_3O_4 Minimum, Manning's.

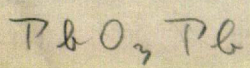
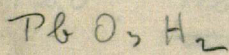
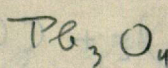
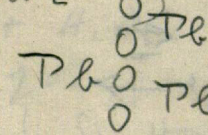
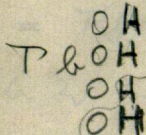
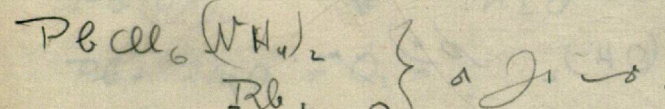
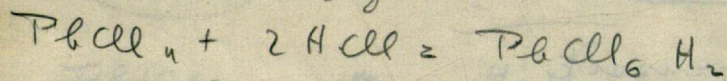
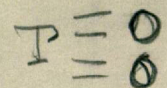


✓ of PbO : PbO & PbO_2 !

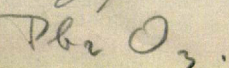
PbO ... $PbCl_2$

PbO_2 ... $PbCl_4$

✓ $PbCl_2$ & $PbCl_4$



(very rare)



Vorlesung

Tl

TlCl

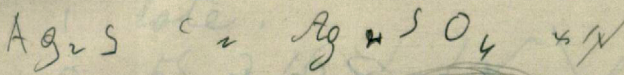
TlOH

TlCl₃

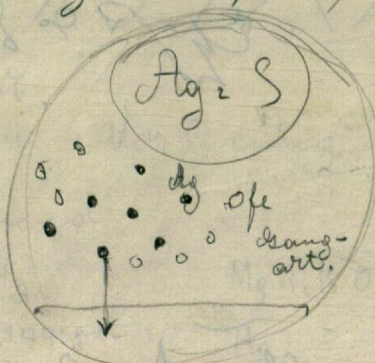
Tl(OH)₃

AgCl

Ag₂S

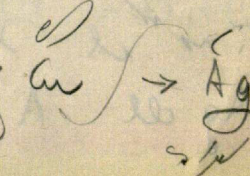
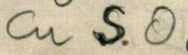
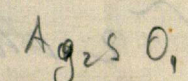
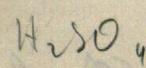


→ AgCl
fe

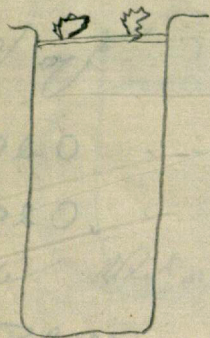


Decomposition

Ag Cu Au



Tyngsboro Det Vilberd. 1 June 68



24/2 0.05.

$$\text{AgNO}_3, \text{AgCl}, \text{Ag}_2\text{O}$$

~ sol. in 22 AgO₂ (Cp₂Co
g sol.)

St. Louis, Mo. Feb 2 1867

Ag dl Ag Br Ag ~~dl~~ 200 g

W. P. B. O. S. N.

Aug Вн

[illegible]

900% δ $\sim$ δ γ

1 Mark H. R.

16 Late.

55 3 16 ✓ ✓ ✓ 1 2 3 4

116 $\mathcal{M}$ 84.

Verleuning

Lilberhydrascyd

Mg Zn Cd Hg

24 65.2 11.18 199.8

1.4 7.1 8.6 13.4

IR. 700° 420° 320° - 30°

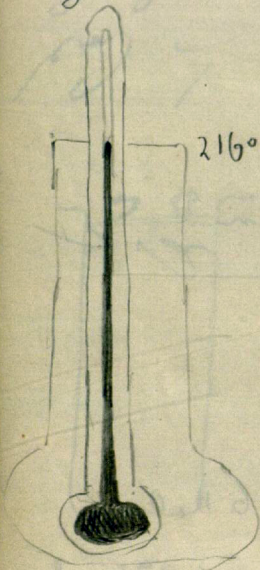
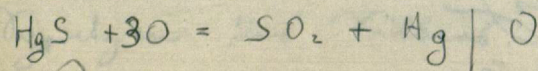
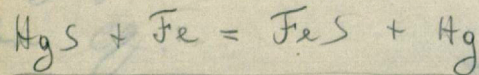
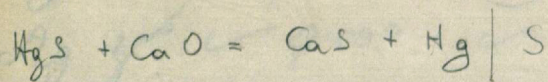
950° 770° 360°

$\text{XIII}_{21} 151 \underline{0} \underline{c} \quad 97 \underline{c} \quad 93 \underline{c} \quad 54 \underline{c}$

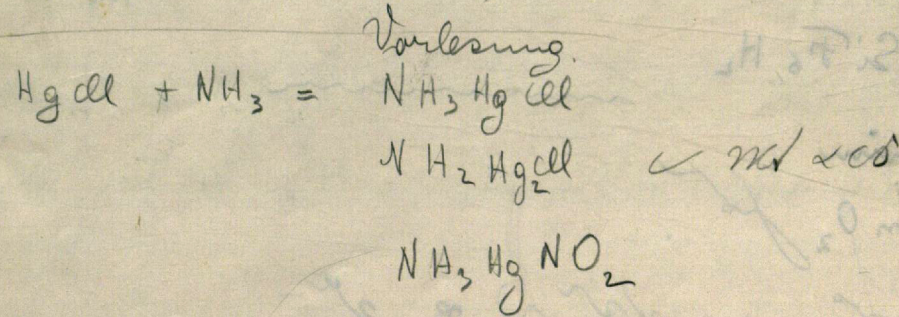
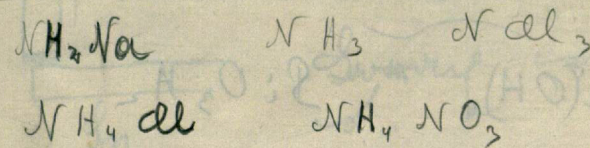
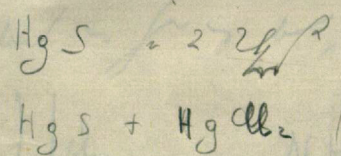
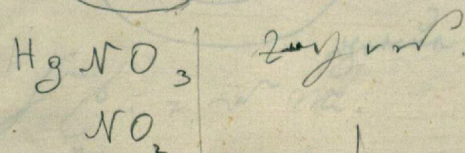
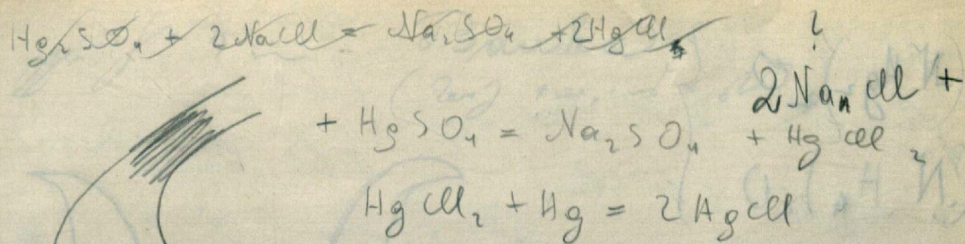
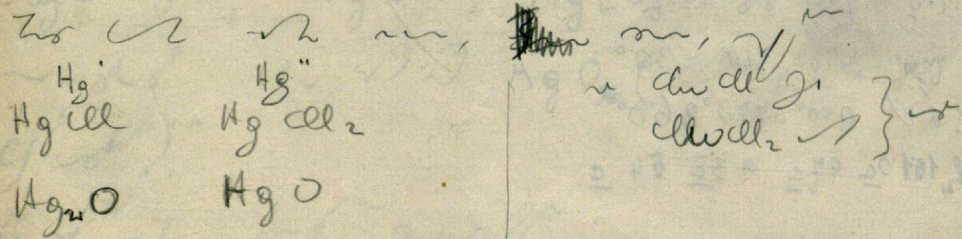
$$\text{MgK}_2(\text{SO}_4)_2 + 6\text{H}_2\text{O}$$

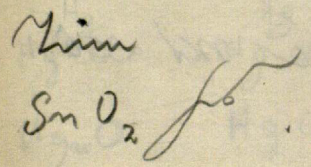
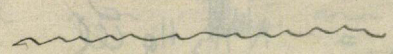
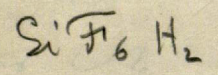
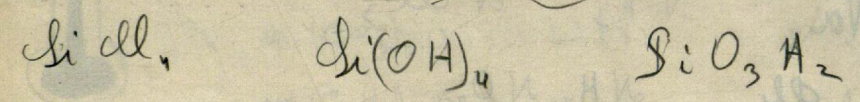
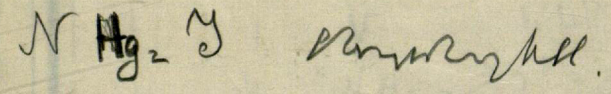
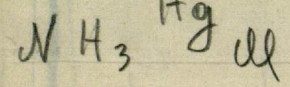
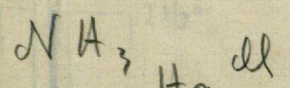
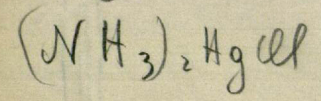
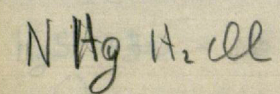
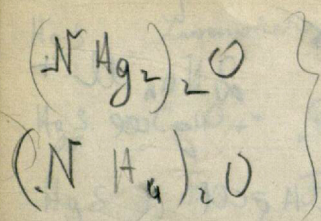
179

HgS Zinnmer.



	Cr	
	Cr	1.66
1	O	1.40
1	N	1.41
4	NH_3	1.30
9	$(\text{CH}_3)_2\text{O}$	1.11
15	$(\text{C}_2\text{H}_5)_2\text{O}$	1.02

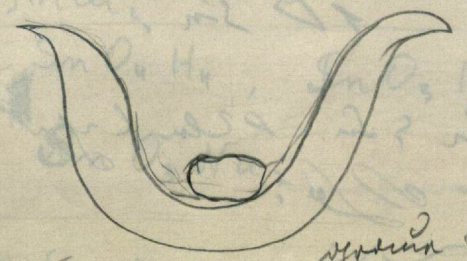




für, auf der 2. u.

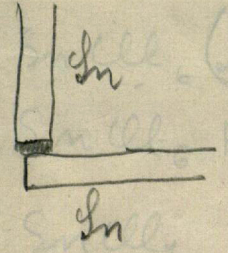


Banka
(200 m, 1. u. 2. u.)



grosse Muldenbildung.
für 2. u. 3. u.

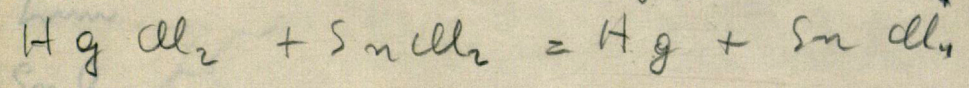
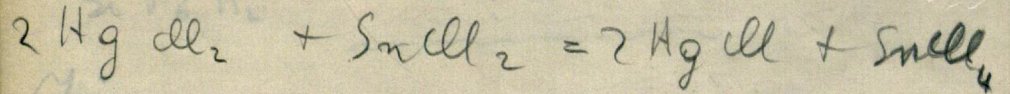
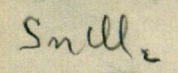
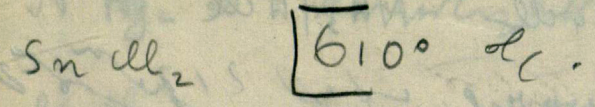
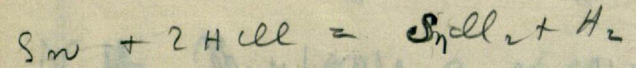
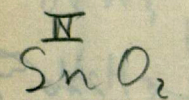
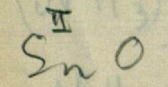
ganz so, wie es ist, für 2. u. 3. u.



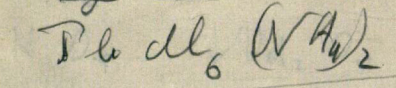
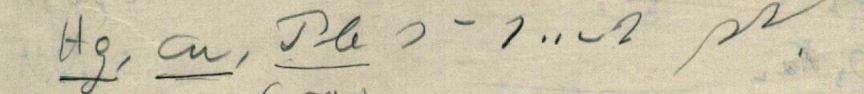
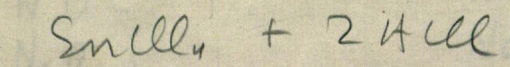
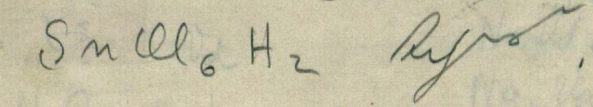
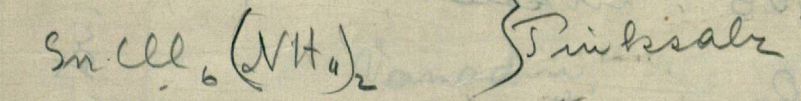
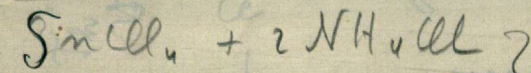
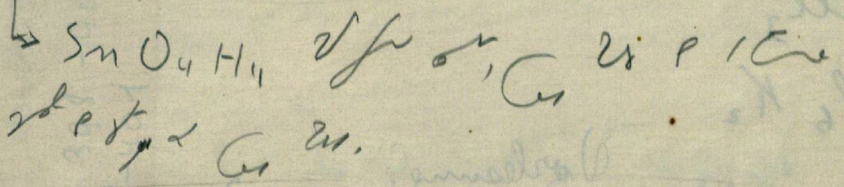
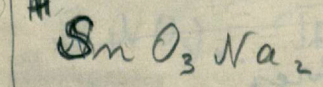
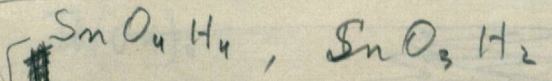
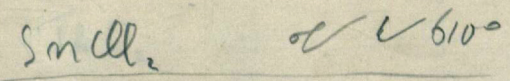
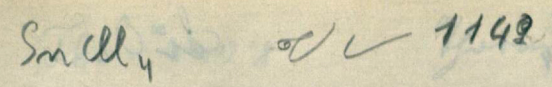
Sammlung von 21. u. 22. u.

$\text{V} \quad 90^\circ \text{ Cu} \quad \text{Sn}$
 $\text{V} \quad 2^\circ \text{ Cu} \quad \text{Sn}$
 $\text{Cu} \quad 4^\circ \text{ Cu} \quad \text{Sn}$

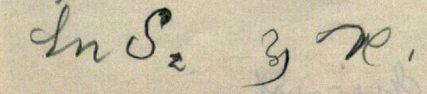
Cryst. of Sn & Cu & Sn . O_2 Cu Cryst
 of Sn . O_2 Cu Cryst .

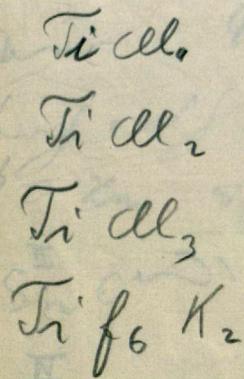
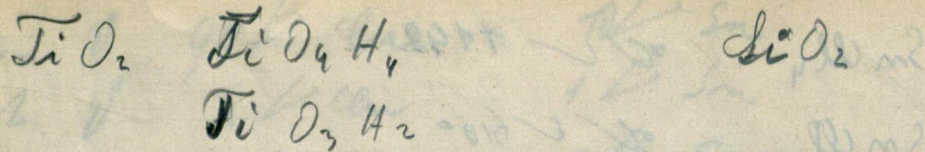


SnCl_4 Cu Cryst
 SnCl_4 Cu Cryst



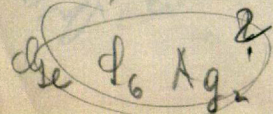
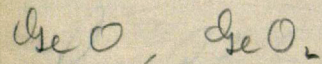
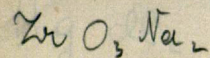
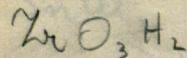
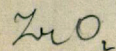
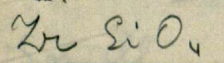
Sn



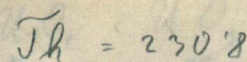


~~Linungeschrei~~

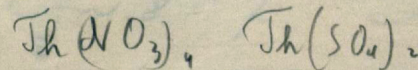
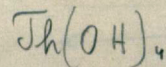
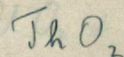
Vorlesung.



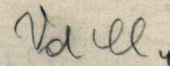
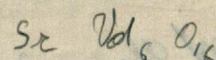
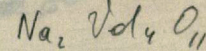
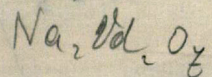
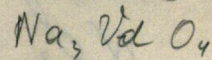
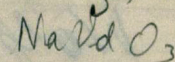
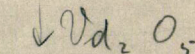
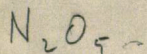
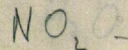
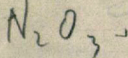
Se. ? dringend

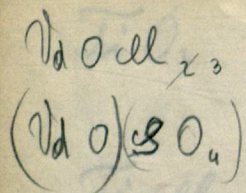


Thorium

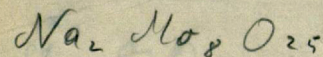
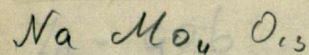
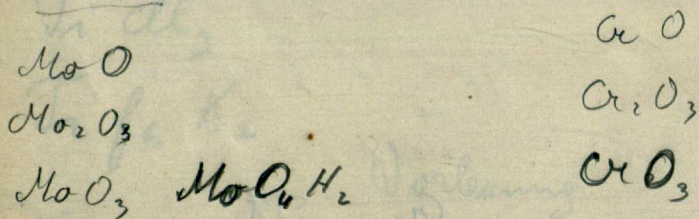


Vanadin

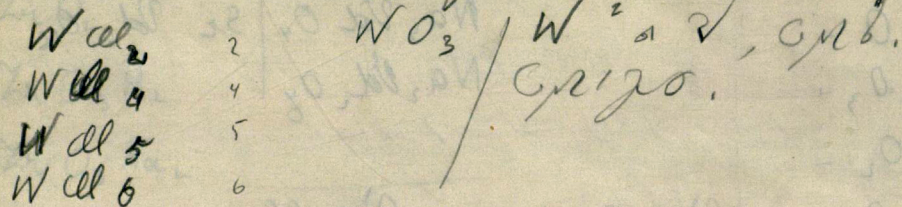




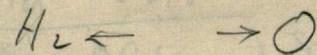
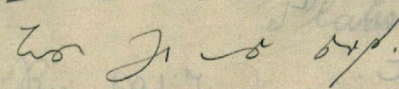
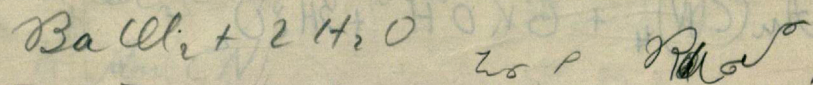
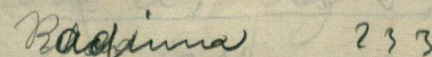
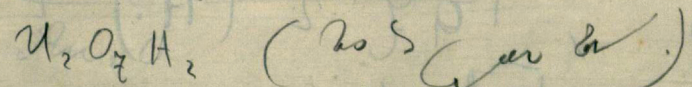
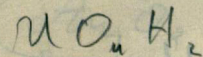
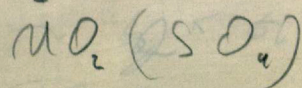
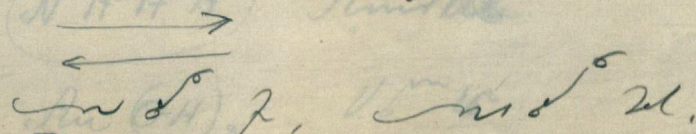
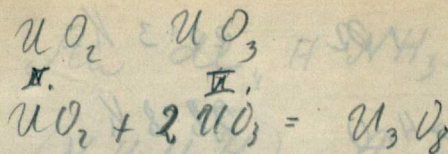
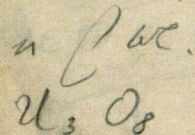
Mo, Cr



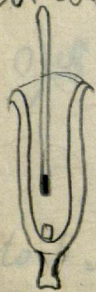
Wolfram



Uran



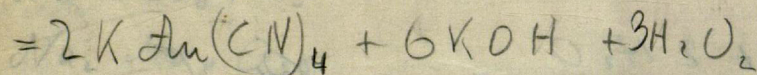
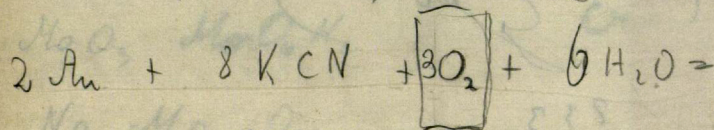
Rd dl,



233g 12 22.3 K.
1g $\rightarrow$ H₂O = 28.8 K

Vorlesung.

An 196.6, 197.2 (H.)



24 Karat = 1 Mark

58.3% = 14 Karat.

An dl

An dl₃ ~~25.5~~

An dl₄ H An dl₄ Na (Lie)

An dl₄ H NH₃

N H H H Amide.

An (OH)₃ 1/6^m 16

$\sim \frac{1}{2} \frac{1}{2} \frac{1}{2}$

An O₂ H = f R. gelblich

An₂O

An dl, Ag dl, Cu dl, Zn

K₂Cu₂(CN)₄ oder K₂Cu(Cy)₂ ?

K₂An(CN)₂

Platinmetalle:

Ru 101.7

Rh 103.0

Pd 106.5

Os 191

Ir 193

Pt 194.8

11-13 / 11

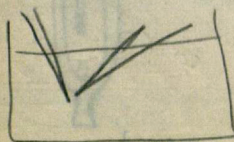
5 22

OsO₄

RuO₄

$$\text{Pt dl}_6 (\text{NH}_4)_2$$

in vol 2 of *Notes on*



Tall

$$Pt \text{ cell, } H_2$$

Pt all,

$$\text{Pt} \text{ d}_6 \text{ H}_2$$
$$\text{Pt}(\text{OH})_4$$

9 R.

$$\text{TAO}_3\text{H}_2$$

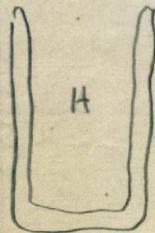
PAO₃Ba 3. 4.

$$\text{Pt}(\text{CN})_4\text{H}_2$$

by John Smith 21-2-22

$$\text{PtCl}_4(\text{NH}_3)_2$$
$$\text{PtCl}_2(\text{NH}_3)_4$$
$$\text{Pd}(\text{NO}_3)_2$$

$\text{PdCl}_2 \cdot \text{K}_2$



H. W. L. H. H. H.

22 May 1894.

$$\text{YHCl}_2, \text{YHCl}_4, \text{YHCl}_3$$
$$\text{MnCl}_2 \cdot \text{K}_2$$

四、

 $O_s \quad O_4 \quad K_2$

CH
CHO
CNH
CHON
CHONS
CHONSP[Y]

CO₂ H₂O

Handwritten notes

CaCl₂

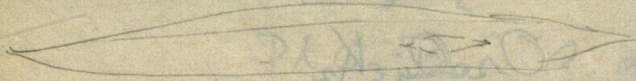
Handwritten notes

Handwritten notes

C N CN

Na

Handwritten notes



Handwritten notes

H₂O | CO₂ N
NOH

$$a + b + c = 100$$

$$a + b + c + d < 100$$

C O C H O N

Handwritten notes

0.301 g Essigsäure.

CO₂ 0.439 g H₂O 0.182 g

12 : 32

2 : 16

44, 12

18, 2

C 39.77
6.71
46.48
O 43.52

H 6.71

x C y H z O

$$x C : y H : z O = 39.77 : 6.71 : 43.52$$

$$x \cdot 12 : y : z \cdot 16$$

$$x : y : z = \frac{39.77}{12} : \frac{6.71}{1} : \frac{43.52}{16}$$

C H O
3.33 6.71 3.34

(C₂H₄O)_n 30
C₂H₄O₂ 60
C₃H₆O₃ 90
C₄H₈O₄ 120

2. Essigsäure.

Willkürlich 0.405 g

Ag 0.261 g.

64.4%

$$100 : 64.4 = x : 108$$

C₂H₄O₂ C₄H₈O₄
C₂H₃O₂ Ag C₄H₆O₄ Ag₂

$$x = 167$$

$$-108$$

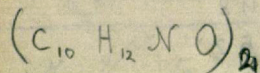
$$+1$$

$$60$$

Chinin

C	H	N	O
73.8	7.60	8.91	9.69
12	1	14	16

6.15	7.60	0.63	0.65
10	12	1	0.95



120
12
84
16
102

de
17%
9.8%

$100 : 9.8 = 10.35 : 1$
 $n = 360$
 35.5
 324.5

$CH=14$ Benzol C_6H_5Br Brombenzol

C_2H_2	$C_{12}H_{12} \dots C_{12}H_{10}Br_2$
C_3H_4	
C_4H_6	
C_5H_8	
C_6H_{10}	

$C_2H_4O_2$ Essigsäure C_2H_5OH Ethanol

CH_2O Formaldehyd

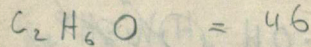
$C_3H_6O_3$

$C_6H_{12}O_6$

Dextrose $C_6H_{12}O_6$

Benzol C_6H_6

Alkohol (C_2H_5O)



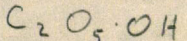
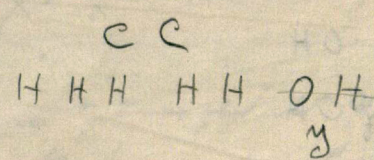
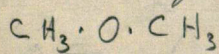
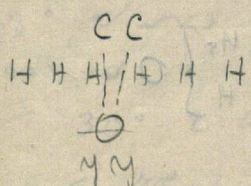
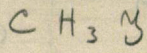
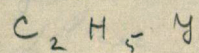
Dimethylether (20)

Ysomerie / C_2H_5ONa

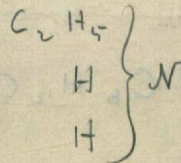
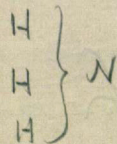
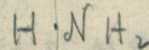
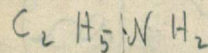
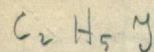
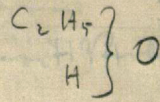
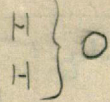
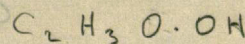
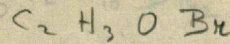
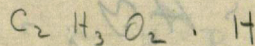
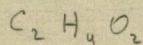
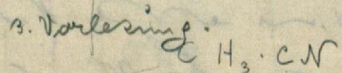
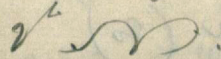
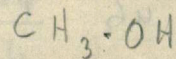
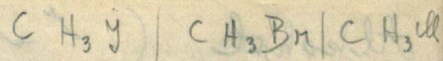
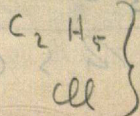
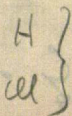
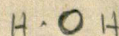
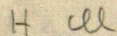
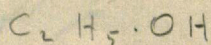
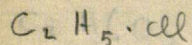
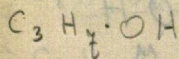
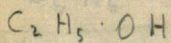
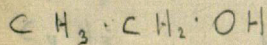
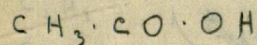
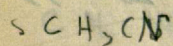
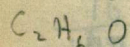
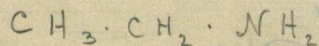
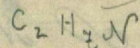
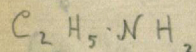
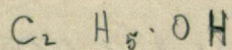
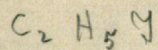
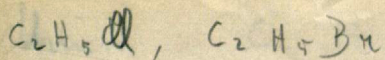
$C \sim 2 \dots Na$
 $2 \dots Na$

CH_3

$CH_3 \dots H$
 $CH_3 \dots H$



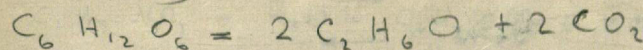
$20 \dots$



Alkohol

2nd. Revised (R 3 of Nov.).

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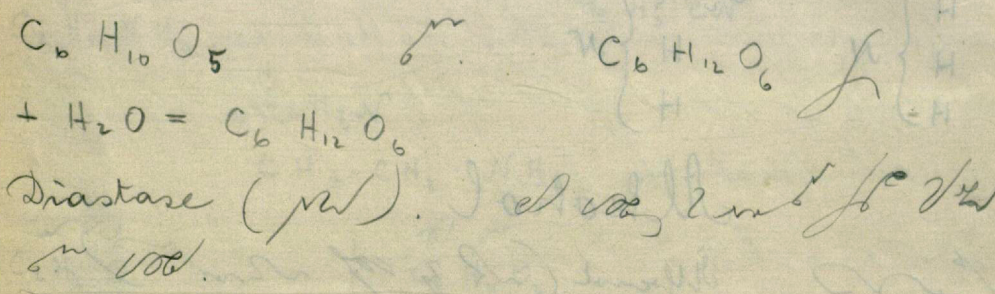
Zymare (51 st wt).

26 we got, and so

✓ 78° dms } mpt g
✓ 100° s

✓ 30° 26'. Pyrene - in the W. of.
Hg & Cu 25 m. 2 1/2 y. No. - 20 1/2 y.
2 1/2 y. of 1/2 y. No. 20 1/2 y. - 1/2 y.
2 1/2 y.

$\rho \approx 1.2$ no CH_3
 no CH_3 & CH_2 in CH_3 & CH_2 by
 $\text{C} \approx 5-10\%$ $\text{H} \approx 12\%$
 $\text{N} \approx 50\%$ $\text{O} \approx 24\%$ $\text{H} \approx 3-5\%$



Vol.	Gen.
95%	92.5%
80 "	73.7 "
5 "	5.08 "

$0.850, 0.92$
 Alcohol
 Wasser
 $\text{H}_2\text{SO}_4, \text{KAl} / ; < \text{CaO}, \text{BaO}; \text{KVO}^2\text{Co}.$

