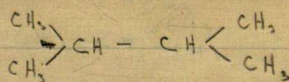
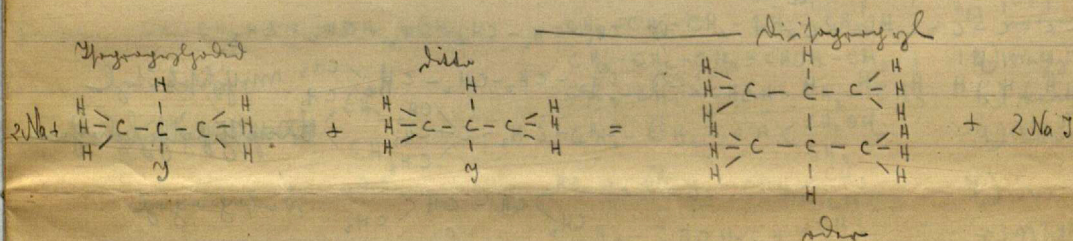
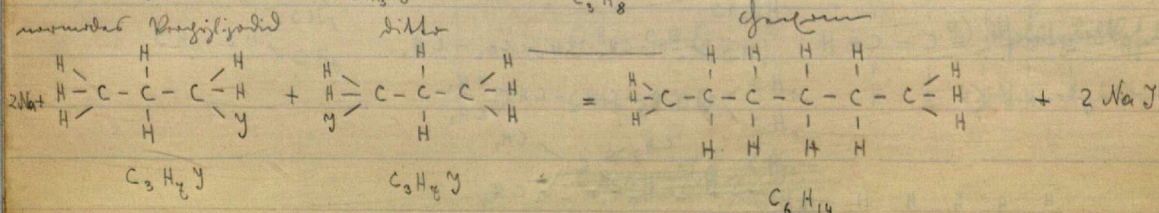
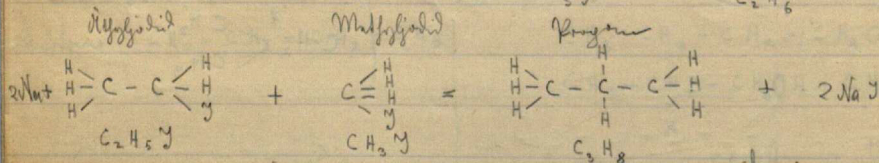
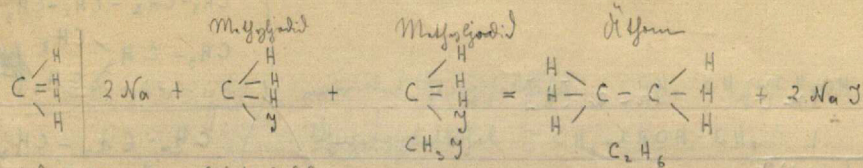
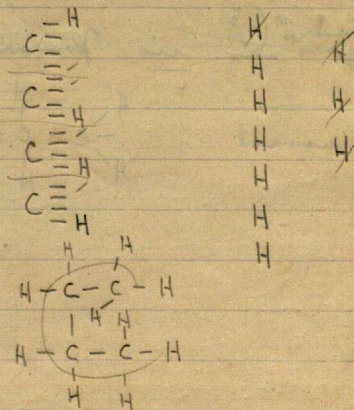
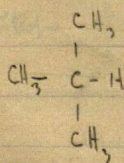
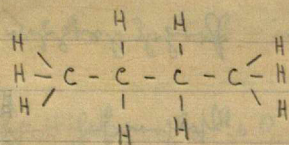
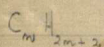


Organische Chemie.
(Holleman)

I.

L. L. C. C_4H_{10} 



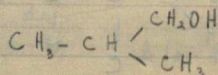
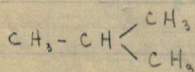
Propan: $\text{CH}_3 - \text{CH}_2 - \text{CH}_3$; normaler Propylalkohol: $\text{CH}_3 - \text{CH}_2 - \text{CH}_2\text{OH}$

Isopropylalkohol: $\text{CH}_3 - \text{CHOH} - \text{CH}_3$

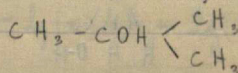
Butan: $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_3$

$\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2\text{OH}$ 1.) normal primär

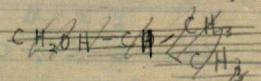
$\text{CH}_3 - \text{CH}_2 - \text{CH}(\text{OH}) - \text{CH}_3$ 2.) normal sekundär



3.) isopropylalkohol



4.) isopropylalkohol



Pentan: $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{CH}_3$

$\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2\text{OH}$

1.) normal primär

$\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CHOH} - \text{CH}_3$

4.) methylpropylalkohol

$\text{CH}_3 - \text{CH}_2 - \text{CHOH} - \text{CH}_2 - \text{CH}_3$

6.) isopropylalkohol

$\text{CH}_3 - \text{CH}_2 - \text{CH} \begin{matrix} \text{CH}_2\text{OH} \\ \text{CH}_3 \end{matrix}$

3.) isopropylalkohol

$\text{CH}_3 - \text{CH}_2 - \text{COH} \begin{matrix} \text{CH}_3 \\ \text{CH}_3 \end{matrix}$

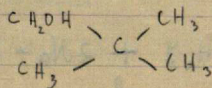
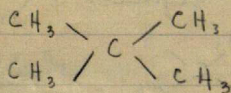
4.) isopropylalkohol

$\text{CH}_3 - \text{CHOH} - \text{CH} \begin{matrix} \text{CH}_3 \\ \text{CH}_3 \end{matrix}$

5.) methylpropylalkohol

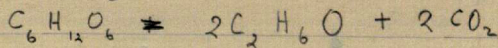
$\text{CH}_2\text{OH} - \text{CH}_2 - \text{CH} \begin{matrix} \text{CH}_3 \\ \text{CH}_3 \end{matrix}$

2.) isopropylalkohol



8.) isopropylalkohol

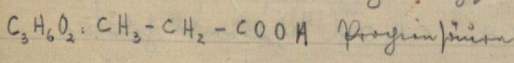
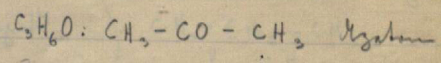
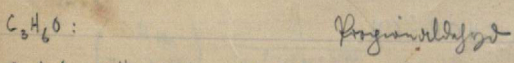
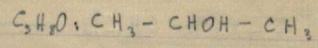
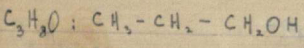
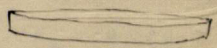
Glukose: $\text{C}_6\text{H}_{12}\text{O}_6$



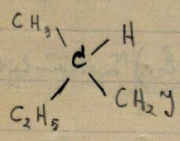
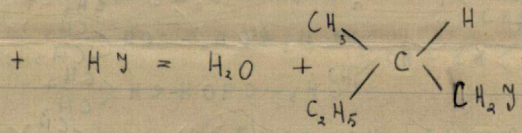
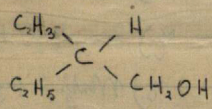
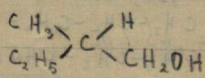
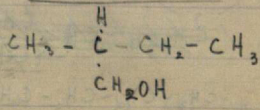
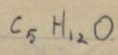
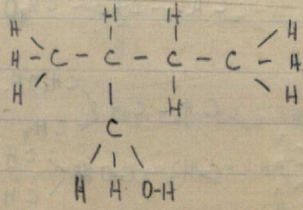
Glukose

+ 2 3 4 5 6 7 8 9 0

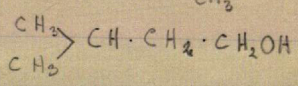
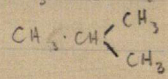
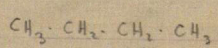
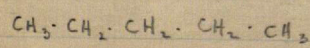
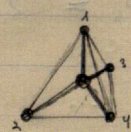
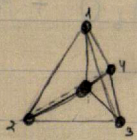
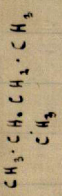
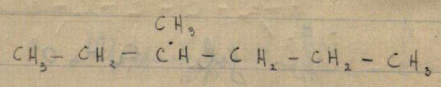
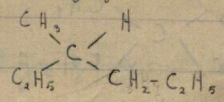
Proyektionsfol:



Valenz der Kohlenstoffverbindungen:



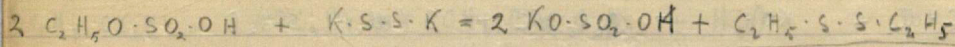
im folgenden



C	H	O
m	$2m + \text{X}$	1
n	$2n + \text{X}$	*
C_{m+n}	$H_{2(m+n)+2}$	0 + H_2O

II.

Mittel Differenz

Markenglas: $\text{C}_2\text{H}_5\text{SH}$ Yissillfar: $\text{C}_2\text{H}_5 \cdot \text{S} \cdot \text{C}_2\text{H}_5$

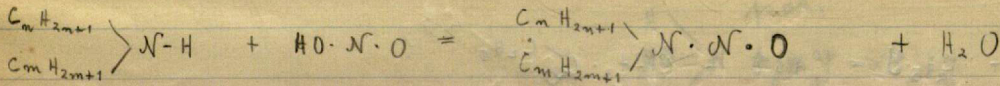
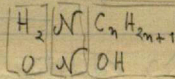
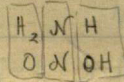
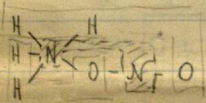
$$\frac{dx}{dt} = k(a-x)(b-x)$$

$$dt = \frac{1}{k} \int \frac{dx}{(a-x)(b-x)} = \frac{1}{k(a-b)} \int \frac{dx}{(b-x)} - \frac{1}{k(a-b)} \int \frac{dx}{(a-x)}$$

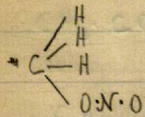
$$t = \frac{1}{k(a-b)} (-\lg(x-b) + \lg(x-a))$$

$$t = \frac{1}{k(a-b)} \lg \frac{a-x}{b-x}$$

$$k = \frac{\lg \frac{a-x}{b-x}}{t(a-b)} \quad a, b$$

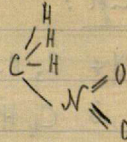


Nukleofide:

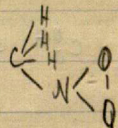


Methylamin

und

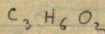
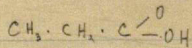


Methylamin

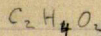
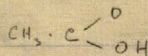


Methylamin

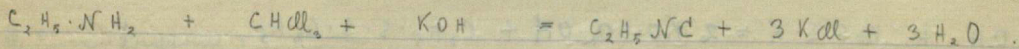
Benzolamine



Häufigkeit:

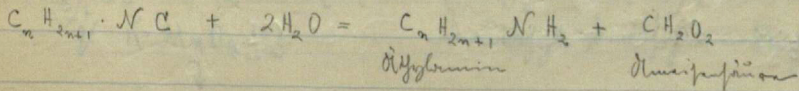


Quercyglucosamine : $C_n H_{2n+1} - N^{\circ} \equiv C$

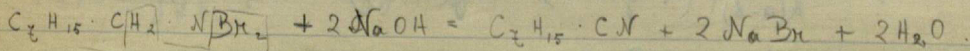
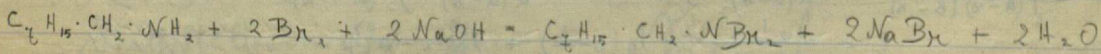


Banknoten auf primitive America

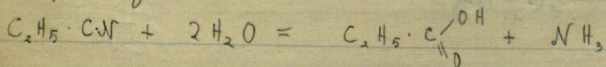
Geophyten der Borealanne tief Tundra



Mittele:

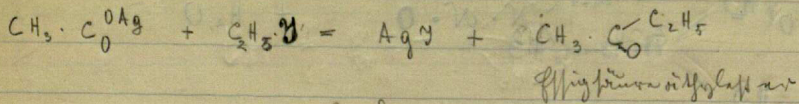


Verfeinerung der Nitrola

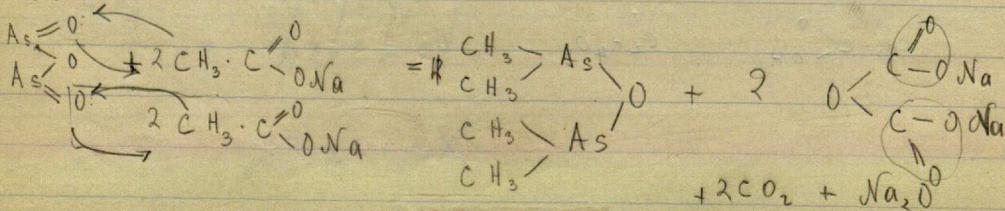
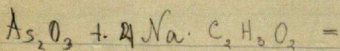
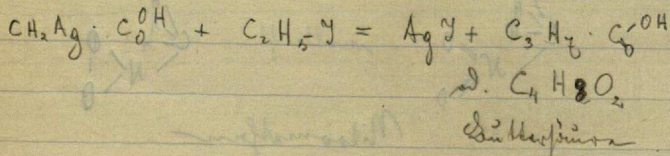


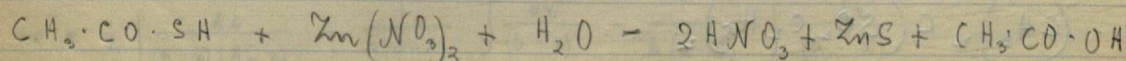
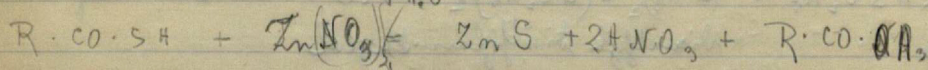
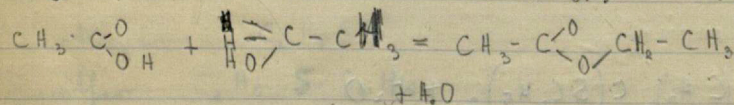
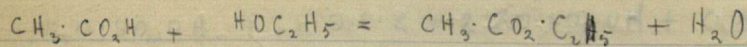
alkylieren: $C_n H_{2n} O_2$

Barbonylgruppe: $-C(=O)-OH$



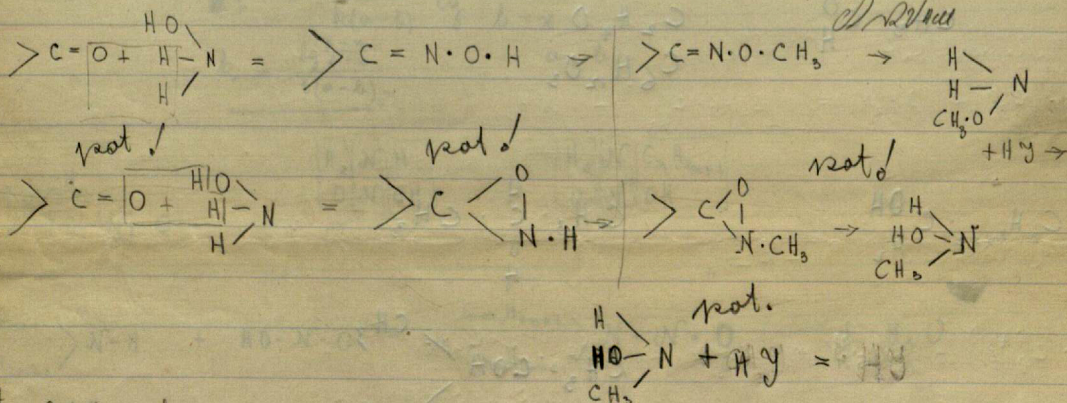
und nicht:



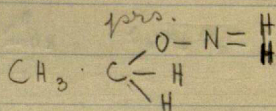
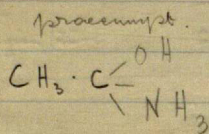
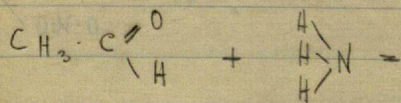
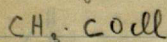


Yisaffig pira

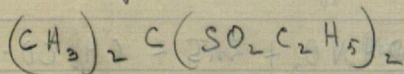
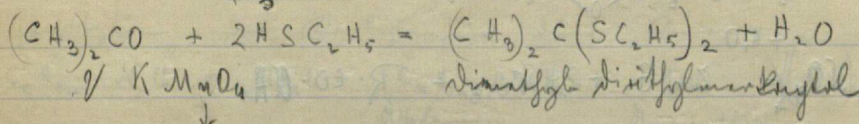
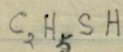
affig pira



Azetylchlorid.



Äthylmercaptan



Dimethyläthylsulfon = Thionol.

